

Astrobiology An Overview

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 Graduate School 😊

November 20-24, 2023

Daily: 10:00-12:00 & **12:45-13:30**



Astrobiology

An Overview

Day 5



Search for Extraterrestrial Intelligence; Alien Biochemistry

<https://www.cosmos.esa.int/web/astrobio/imprs-2023>

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|--------------------------|--|
| Monday November 20 | Day 1: Definition of Life; Origin of Life; Evolution of Life; Limits of Life 10:00-12:00 & 13:00-14:00 |
| Tuesday November 21 | Day 2: Earth Climate History; Mars and Venus Climates 10:00-12:00 & 13:00-14:00 OLD SEMINAR ROOM |
| Wednesday November 22 | Day 3: Habitable Places in the Solar System; Mars; Moons of Giant Planets 10:00-12:00 & 13:00-14:00 |
| Thursday November 23 | Day 4: Habitable Places beyond the Solar System; Exoplanets properties; Biosignatures 10:00-12:00 & 13:00-14:00 |
| Friday November 24 | Day 5: Search for Extraterrestrial Intelligence; Alien Biochemistry 10:00-12:00 & 12:45-13:30 |

Alien Biochemistries



Psychrolutes microporos AMS I.42771-001



What should we be looking for
when we are looking for
“Life” on exoplanets?

Let's consider Life as:

Complex chemical reactions that allow

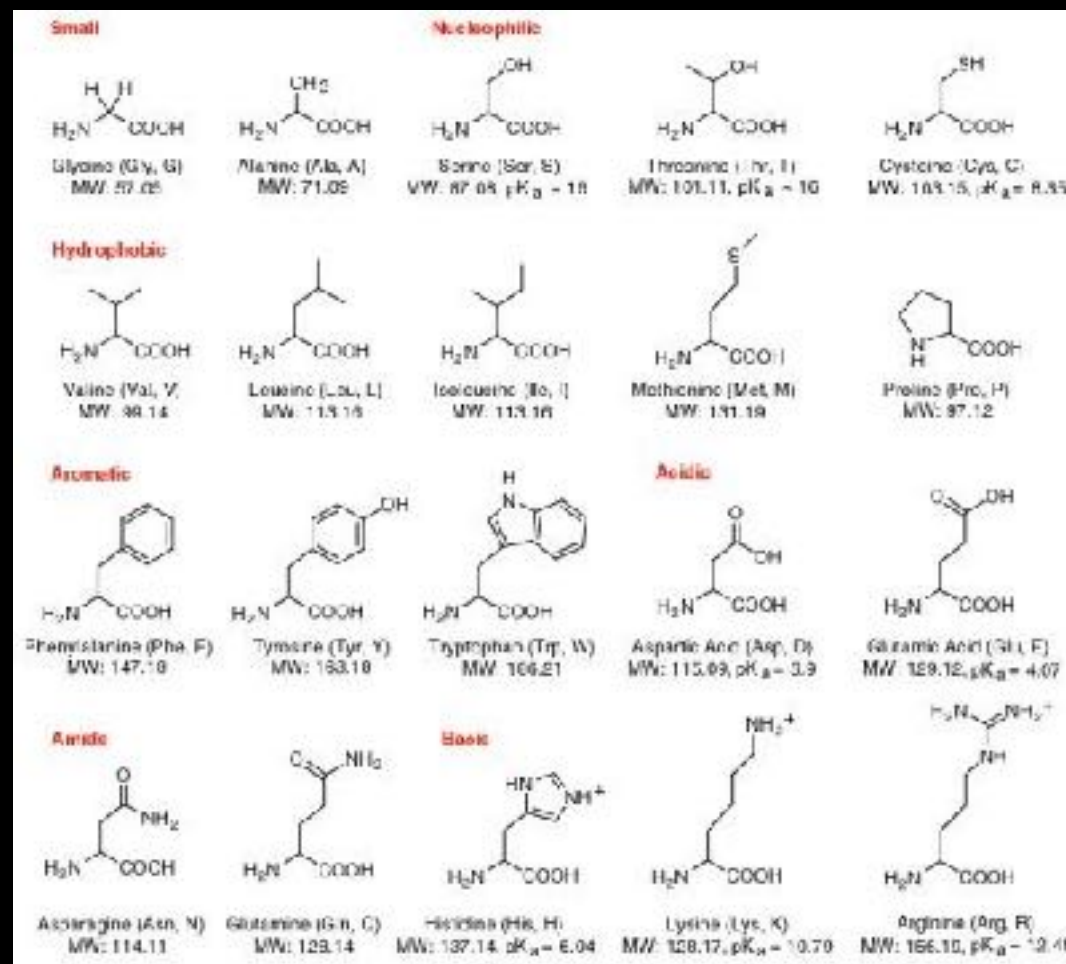
- Metabolism
- Reproduction
- Evolution

Different Biopolymers

Different Amino acids

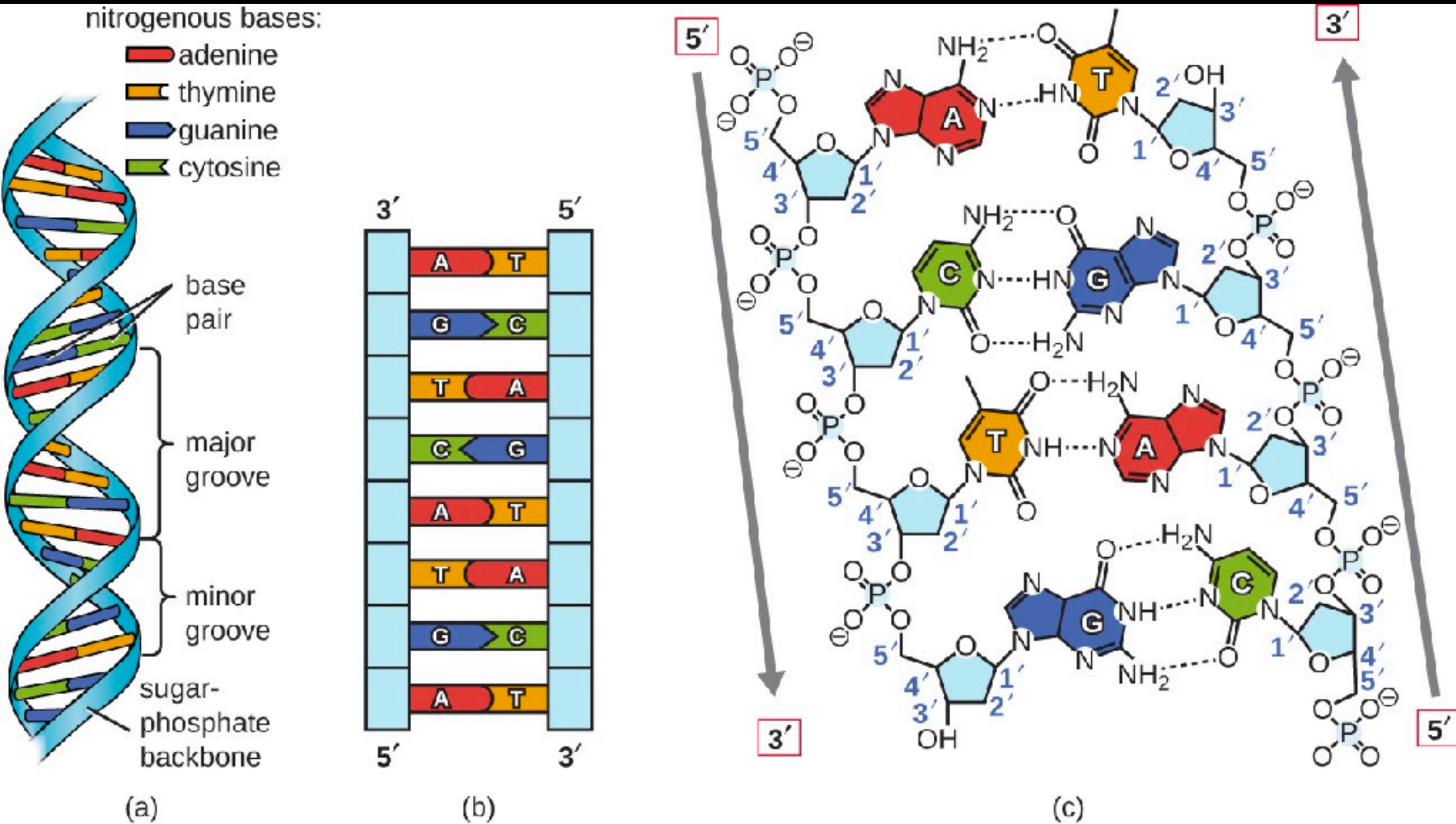
All life on Earth uses the standard 20 amino acids
(these are not the most abundant ones found in Meteorites)

Ribosomes are capable of using different amino acids to synthesise proteins



DNA-based life might exist, similar in all other respect to Earth life,
but using different amino acids to synthesize proteins

Different “DNA”



Different “DNA”

Base-pairing is not unique!

Large purines (A,G) pair with small pyrimidine (C,T,U) - other combinations are possible...
(and have been synthesized in the lab)

The **number of “letters”** (3) to code an amino acid could be different (e.g $4^2=16$, $4^4=256$ AA be coded). Life could build on a different set of >20 ‘standard’ amino acids

The **Ribose backbone** could be replaced (e.g. by glycerol). See the pre-RNA world research...

The **Phosphor Group** in the backbone could be modified. e.g. P replaced by Ar - claims for ‘Arsenic life’ were made (but refuted)



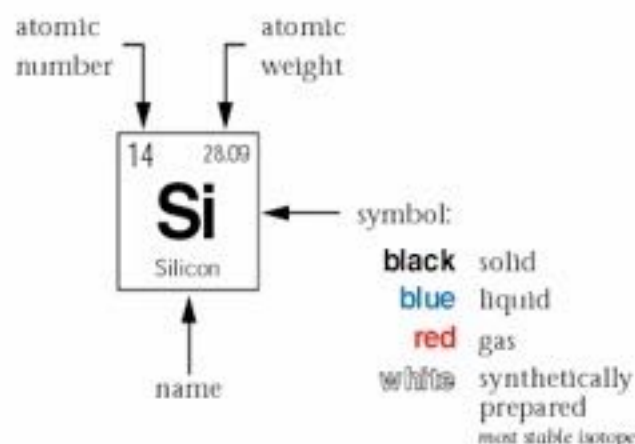
Periodic Table of the Elements



ERNEST ORLANDO LAWRENCE
BERKELEY NATIONAL LABORATORY

Skun3 Seaborg

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| <div><div><div>atomic number</div><div>atomic weight</div></div><div><div>14</div><div>28.09</div><div>Si</div><div>Silicon</div></div><div><div>symbol:</div><div>black solid</div><div>blue liquid</div><div>red gas</div><div>white synthetically prepared</div><div>most stable isotope</div></div><div><div>name</div></div></div> <div><div>alkali metals</div><div>alkaline earth metals</div><div>transitional metals</div><div>other metals</div><div>nonmetals</div><div>noble gases</div></div> <div><div>Heun's Seaborg</div></div> | | | | | | | | | | | | | | | | | | <table><tr><td>1</td><td>1.01</td><td colspan="16"></td><td>2</td><td>4.003</td></tr><tr><td>H</td><td>Hydrogen</td><td colspan="16"></td><td>He</td><td>Helium</td></tr><tr><td>3</td><td>6.94</td><td>4</td><td>9.01</td><td colspan="14"></td><td>10</td><td>20.18</td></tr><tr><td>Li</td><td>Lithium</td><td>Be</td><td>Beryllium</td><td colspan="14"></td><td>Ne</td><td>Neon</td></tr><tr><td>11</td><td>22.99</td><td>12</td><td>24.31</td><td colspan="14"></td><td>18</td><td>39.95</td></tr><tr><td>Na</td><td>Sodium</td><td>Mg</td><td>Magnesium</td><td colspan="14"></td><td>Ar</td><td>Argon</td></tr><tr><td>19</td><td>39.10</td><td>20</td><td>40.08</td><td>21</td><td>44.96</td><td>22</td><td>47.90</td><td>23</td><td>50.94</td><td>24</td><td>51.996</td><td>25</td><td>54.94</td><td>26</td><td>55.85</td><td>27</td><td>58.93</td><td>28</td><td>58.70</td><td>29</td><td>63.55</td><td>30</td><td>65.37</td><td>31</td><td>69.72</td><td>32</td><td>72.64</td><td>33</td><td>74.92</td><td>34</td><td>78.96</td><td>35</td><td>79.90</td><td>36</td><td>83.80</td></tr><tr><td>K</td><td>Potassium</td><td>Ca</td><td>Calcium</td><td>Sc</td><td>Scandium</td><td>Ti</td><td>Titanium</td><td>V</td><td>Vanadium</td><td>Cr</td><td>Chromium</td><td>Mn</td><td>Manganese</td><td>Fe</td><td>Iron</td><td>Co</td><td>Cobalt</td><td>Ni</td><td>Nickel</td><td>Cu</td><td>Copper</td><td>Zn</td><td>Zinc</td><td>Ga</td><td>Gallium</td><td>Ge</td><td>Germanium</td><td>As</td><td>Arsenic</td><td>Se</td><td>Selenium</td><td>Br</td><td>Bromine</td><td>Kr</td><td>Krypton</td></tr><tr><td>37</td><td>85.47</td><td>38</td><td>87.62</td><td>39</td><td>88.91</td><td>40</td><td>91.22</td><td>41</td><td>92.91</td><td>42</td><td>95.94</td><td>43</td><td>(98)</td><td>44</td><td>101.07</td><td>45</td><td>102.91</td><td>46</td><td>106.40</td><td>47</td><td>107.87</td><td>48</td><td>112.41</td><td>49</td><td>114.82</td><td>50</td><td>118.69</td><td>51</td><td>121.75</td><td>52</td><td>127.60</td><td>53</td><td>126.90</td><td>54</td><td>131.30</td></tr><tr><td>Rb</td><td>Rubidium</td><td>Sr</td><td>Strontium</td><td>Y</td><td>Yttrium</td><td>Zr</td><td>Zirconium</td><td>Nb</td><td>Niobium</td><td>Mo</td><td>Molybdenum</td><td>Tc</td><td>Technetium</td><td>Ru</td><td>Ruthenium</td><td>Rh</td><td>Rhodium</td><td>Pd</td><td>Palladium</td><td>Ag</td><td>Silver</td><td>Cd</td><td>Cadmium</td><td>In</td><td>Indium</td><td>Sn</td><td>Tin</td><td>Sb</td><td>Antimony</td><td>Te</td><td>Tellurium</td><td>I</td><td>Iodine</td><td>Xe</td><td>Xenon</td></tr><tr><td>55</td><td>132.91</td><td>56</td><td>137.33</td><td>57</td><td>138.91</td><td>72</td><td>178.49</td><td>73</td><td>180.95</td><td>74</td><td>183.85</td><td>75</td><td>186.21</td><td>76</td><td>190.20</td><td>77</td><td>192.22</td><td>78</td><td>196.09</td><td>79</td><td>196.97</td><td>80</td><td>200.59</td><td>81</td><td>204.37</td><td>82</td><td>207.19</td><td>83</td><td>208.98</td><td>84</td><td>(209)</td><td>85</td><td>(210)</td><td>86</td><td>(222)</td></tr><tr><td>Cs</td><td>Cesium</td><td>Ba</td><td>Barium</td><td>La</td><td>Lanthanum</td><td>Hf</td><td>Hafnium</td><td>Ta</td><td>Tantalum</td><td>W</td><td>Tungsten</td><td>Re</td><td>Rhenium</td><td>Os</td><td>Osmium</td><td>Ir</td><td>Iridium</td><td>Pt</td><td>Platinum</td><td>Au</td><td>Gold</td><td>Hg</td><td>Mercury</td><td>Tl</td><td>Thallium</td><td>Pb</td><td>Lead</td><td>Bi</td><td>Bismuth</td><td>Po</td><td>Polonium</td><td>At</td><td>Astatine</td><td>Rn</td><td>Radon</td></tr><tr><td>87</td><td>(223)</td><td>88</td><td>226.03</td><td>89</td><td>227.03</td><td>104</td><td>(261)</td><td>105</td><td>(262)</td><td>106</td><td>(266)</td><td>107</td><td>(262)</td><td>108</td><td>(265)</td><td>109</td><td>(266)</td><td>110</td><td>(271)</td><td>111</td><td>(272)</td><td>112</td><td>(277)</td><td>(113)</td><td>(114)</td><td>(115)</td><td>(116)</td><td>(117)</td><td>(118)</td><td>(119)</td><td>(120)</td><td>(121)</td><td>(122)</td></tr><tr><td>Fr</td><td>Francium</td><td>Ra</td><td>Radium</td><td>Ac</td><td>Actinium</td><td>Rf</td><td>Rutherfordium</td><td>Ha</td><td>Hahnium</td><td>Sg</td><td>Seaborgium</td><td>Bh</td><td>Bohrium</td><td>Hs</td><td>Hassium</td><td>Mt</td><td>Meitnerium</td><td colspan="10"></td></tr></table> | | | | | | | | | | | | | | | | | | 1 | 1.01 | | | | | | | | | | | | | | | | | 2 | 4.003 | H | Hydrogen | | | | | | | | | | | | | | | | | He | Helium | 3 | 6.94 | 4 | 9.01 | | | | | | | | | | | | | | | 10 | 20.18 | Li | Lithium | Be | Beryllium | | | | | | | | | | | | | | | Ne | Neon | 11 | 22.99 | 12 | 24.31 | | | | | | | | | | | | | | | 18 | 39.95 | Na | Sodium | Mg | Magnesium | | | | | | | | | | | | | | | Ar | Argon | 19 | 39.10 | 20 | 40.08 | 21 | 44.96 | 22 | 47.90 | 23 | 50.94 | 24 | 51.996 | 25 | 54.94 | 26 | 55.85 | 27 | 58.93 | 28 | 58.70 | 29 | 63.55 | 30 | 65.37 | 31 | 69.72 | 32 | 72.64 | 33 | 74.92 | 34 | 78.96 | 35 | 79.90 | 36 | 83.80 | K | Potassium | Ca | Calcium | Sc | Scandium | Ti | Titanium | V | Vanadium | Cr | Chromium | Mn | Manganese | Fe | Iron | Co | Cobalt | Ni | Nickel | Cu | Copper | Zn | Zinc | Ga | Gallium | Ge | Germanium | As | Arsenic | Se | Selenium | Br | Bromine | Kr | Krypton | 37 | 85.47 | 38 | 87.62 | 39 | 88.91 | 40 | 91.22 | 41 | 92.91 | 42 | 95.94 | 43 | (98) | 44 | 101.07 | 45 | 102.91 | 46 | 106.40 | 47 | 107.87 | 48 | 112.41 | 49 | 114.82 | 50 | 118.69 | 51 | 121.75 | 52 | 127.60 | 53 | 126.90 | 54 | 131.30 | Rb | Rubidium | Sr | Strontium | Y | Yttrium | Zr | Zirconium | Nb | Niobium | Mo | Molybdenum | Tc | Technetium | Ru | Ruthenium | Rh | Rhodium | Pd | Palladium | Ag | Silver | Cd | Cadmium | In | Indium | Sn | Tin | Sb | Antimony | Te | Tellurium | I | Iodine | Xe | Xenon | 55 | 132.91 | 56 | 137.33 | 57 | 138.91 | 72 | 178.49 | 73 | 180.95 | 74 | 183.85 | 75 | 186.21 | 76 | 190.20 | 77 | 192.22 | 78 | 196.09 | 79 | 196.97 | 80 | 200.59 | 81 | 204.37 | 82 | 207.19 | 83 | 208.98 | 84 | (209) | 85 | (210) | 86 | (222) | Cs | Cesium | Ba | Barium | La | Lanthanum | Hf | Hafnium | Ta | Tantalum | W | Tungsten | Re | Rhenium | Os | Osmium | Ir | Iridium | Pt | Platinum | Au | Gold | Hg | Mercury | Tl | Thallium | Pb | Lead | Bi | Bismuth | Po | Polonium | At | Astatine | Rn | Radon | 87 | (223) | 88 | 226.03 | 89 | 227.03 | 104 | (261) | 105 | (262) | 106 | (266) | 107 | (262) | 108 | (265) | 109 | (266) | 110 | (271) | 111 | (272) | 112 | (277) | (113) | (114) | (115) | (116) | (117) | (118) | (119) | (120) | (121) | (122) | Fr | Francium | Ra | Radium | Ac | Actinium | Rf | Rutherfordium | Ha | Hahnium | Sg | Seaborgium | Bh | Bohrium | Hs | Hassium | Mt | Meitnerium | | | | | | | | | | |
| 1 | 1.01 | | | | | | | | | | | | | | | | | 2 | 4.003 | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| H | Hydrogen | | | | | | | | | | | | | | | | | He | Helium | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 3 | 6.94 | 4 | 9.01 | | | | | | | | | | | | | | | 10 | 20.18 | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| Li | Lithium | Be | Beryllium | | | | | | | | | | | | | | | Ne | Neon | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 11 | 22.99 | 12 | 24.31 | | | | | | | | | | | | | | | 18 | 39.95 | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| Na | Sodium | Mg | Magnesium | | | | | | | | | | | | | | | Ar | Argon | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 19 | 39.10 | 20 | 40.08 | 21 | 44.96 | 22 | 47.90 | 23 | 50.94 | 24 | 51.996 | 25 | 54.94 | 26 | 55.85 | 27 | 58.93 | 28 | 58.70 | 29 | 63.55 | 30 | 65.37 | 31 | 69.72 | 32 | 72.64 | 33 | 74.92 | 34 | 78.96 | 35 | 79.90 | 36 | 83.80 | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| K | Potassium | Ca | Calcium | Sc | Scandium | Ti | Titanium | V | Vanadium | Cr | Chromium | Mn | Manganese | Fe | Iron | Co | Cobalt | Ni | Nickel | Cu | Copper | Zn | Zinc | Ga | Gallium | Ge | Germanium | As | Arsenic | Se | Selenium | Br | Bromine | Kr | Krypton | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 37 | 85.47 | 38 | 87.62 | 39 | 88.91 | 40 | 91.22 | 41 | 92.91 | 42 | 95.94 | 43 | (98) | 44 | 101.07 | 45 | 102.91 | 46 | 106.40 | 47 | 107.87 | 48 | 112.41 | 49 | 114.82 | 50 | 118.69 | 51 | 121.75 | 52 | 127.60 | 53 | 126.90 | 54 | 131.30 | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| Rb | Rubidium | Sr | Strontium | Y | Yttrium | Zr | Zirconium | Nb | Niobium | Mo | Molybdenum | Tc | Technetium | Ru | Ruthenium | Rh | Rhodium | Pd | Palladium | Ag | Silver | Cd | Cadmium | In | Indium | Sn | Tin | Sb | Antimony | Te | Tellurium | I | Iodine | Xe | Xenon | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 55 | 132.91 | 56 | 137.33 | 57 | 138.91 | 72 | 178.49 | 73 | 180.95 | 74 | 183.85 | 75 | 186.21 | 76 | 190.20 | 77 | 192.22 | 78 | 196.09 | 79 | 196.97 | 80 | 200.59 | 81 | 204.37 | 82 | 207.19 | 83 | 208.98 | 84 | (209) | 85 | (210) | 86 | (222) | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| Cs | Cesium | Ba | Barium | La | Lanthanum | Hf | Hafnium | Ta | Tantalum | W | Tungsten | Re | Rhenium | Os | Osmium | Ir | Iridium | Pt | Platinum | Au | Gold | Hg | Mercury | Tl | Thallium | Pb | Lead | Bi | Bismuth | Po | Polonium | At | Astatine | Rn | Radon | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| 87 | (223) | 88 | 226.03 | 89 | 227.03 | 104 | (261) | 105 | (262) | 106 | (266) | 107 | (262) | 108 | (265) | 109 | (266) | 110 | (271) | 111 | (272) | 112 | (277) | (113) | (114) | (115) | (116) | (117) | (118) | (119) | (120) | (121) | (122) | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |
| Fr | Francium | Ra | Radium | Ac | Actinium | Rf | Rutherfordium | Ha | Hahnium | Sg | Seaborgium | Bh | Bohrium | Hs | Hassium | Mt | Meitnerium | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | | |



- alkali metals
- alkaline earth metals
- transitional metals
- other metals
- nonmetals
- noble gases

Lanthanide series ▶

| | | | | | | | | | | | | | |
|--------------------|--------------------------|-----------------------|------------------------|----------------------|----------------------|------------------------|---------------------|------------------------|---------------------|--------------------|---------------------|-----------------------|----------------------|
| 58 Ce Cerium | 59 Pr Praseodymium | 60 Nd Neodymium | 61 Pm Promethium | 62 Sm Samarium | 63 Eu Europium | 64 Gd Gadolinium | 65 Tb Terbium | 66 Dy Dysprosium | 67 Ho Holmium | 68 Er Erbium | 69 Tm Thulium | 70 Yb Ytterbium | 71 Lu Lutetium |
|--------------------|--------------------------|-----------------------|------------------------|----------------------|----------------------|------------------------|---------------------|------------------------|---------------------|--------------------|---------------------|-----------------------|----------------------|

Actinide series ▶

| | | | | | | | | | | | | | |
|---------------------|--------------------------|--------------------|-----------------------|-----------------------|-----------------------|--------------------|-----------------------|-------------------------|-------------------------|----------------------|--------------------------|-----------------------|-------------------------|
| 90 Th Thorium | 91 Pa Protactinium | 92 U Uranium | 93 Np Neptunium | 94 Pu Plutonium | 95 Am Americium | 96 Cm Curium | 97 Bk Berkelium | 98 Cf Californium | 99 Es Einsteinium | 100 Fm Fermium | 101 Md Mendelevium | 102 No Nobelium | 103 Lr Lawrencium |
|---------------------|--------------------------|--------------------|-----------------------|-----------------------|-----------------------|--------------------|-----------------------|-------------------------|-------------------------|----------------------|--------------------------|-----------------------|-------------------------|

Different “DNA”

Arsenic Life: Wolfe-Simon et al. 2010

Really cool - but unfortunately wrong

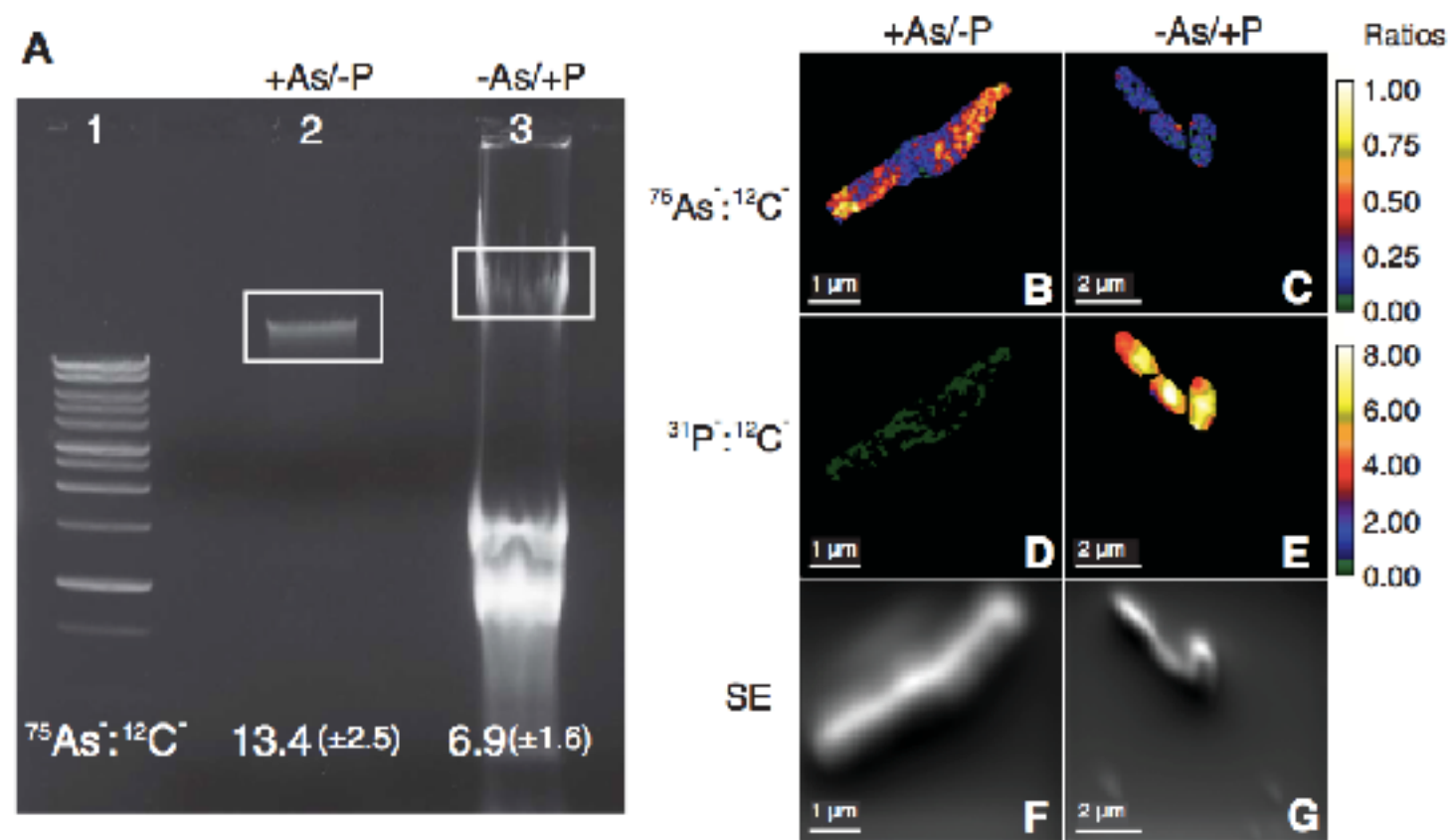


Fig. 2. NanoSIMS analyses of GFAJ-1: extracted DNA and whole-cells elemental ratio maps. (A) Agarose gel loaded with DNA/RNA extracted from GFAJ-1 grown (lane 2) +As/-P and (lane 3) -As/+P as compared with (lane 1) a DNA standard. Genomic bands were excised as indicated and analyzed with NanoSIMS. Ion ratios of $^{75}\text{As}^-:^{12}\text{C}^-$ of excised gel bands are indicated below with 2σ error shown (all values multiplied by 10^{-6}). (B to G) NanoSIMS images of whole GFAJ-1 cells grown either [(B), (D), and (F)] +As/-P or [(C), (E), and (G)] -As/+P. Shown are the ion ratios of [(B) and (C)] $^{75}\text{As}^-:^{12}\text{C}^-$, [(D) and (E)] $^{31}\text{P}^-:^{12}\text{C}^-$, and [(F) and (G)] secondary electron (SE). Ratios in (B) and (C) are multiplied by 10^{-4} and in (D) and (E) are multiplied by 10^{-3} . The color bars indicate measured elemental ratios on a log scale as indicated. Length scale is as indicated on images; images contain equivalent pixel density (11).



RESEARCH ARTICLE

A Bacterium That Can Grow by Using Arsenic Instead of Phosphorus

Felisa Wolfe-Simon,^{1,2*} Jodi Switzer Blum,² Thomas R. Kulp,² Gwyneth W. Gordon,³ Shelley E. Hoefft,² Jennifer Pett-Ridge,⁴ John F. Stolz,⁵ Samuel M. Webb,⁶ Peter K. Weber,⁴ Paul C. W. Davies,^{1,7} Ariel D. Anbar,^{1,3,8} Ronald S. Oremland⁹

Life is mostly composed of the elements carbon, hydrogen, nitrogen, oxygen, sulfur, and phosphorus. Although these six elements make up nucleic acids, proteins, and lipids and thus the bulk of living matter, it is theoretically possible that some other elements in the periodic table could serve the same functions. Here, we describe a bacterium, strain GFAJ-1 of the Halomonadaceae, isolated from Mono Lake, California, that is able to substitute arsenic for phosphorus to sustain its growth. Our data show evidence for arsenate in macromolecules that normally contain phosphate, most notably nucleic acids and proteins. Exchange of one of the major bio-elements may have profound evolutionary and geochemical importance.

Biological dependence on the six major nutrient elements carbon, hydrogen, nitrogen, oxygen, sulfur, and phosphorus (P) is complemented by a selected array of other elements, usually metals or metalloids present in trace quantities that serve critical cellular functions, such as enzyme co-factors (1). There are many cases of these trace elements substituting for one another. A few examples include the substitution of tungsten for molybdenum and cadmium for zinc in some enzyme families (2, 3) and copper for iron as an oxygen-carrier in some arthropods and mollusks (4). In these examples and others, the trace elements that interchange share chemical similarities that facilitate the swap. However, there are no prior reports of substitutions for any of the six major elements essential for life. Here, we present evidence that arsenic can substitute for phosphorus in the biomolecules of a naturally occurring bacterium.

Arsenic (As) is a chemical analog of P, which lies directly below P in the periodic table. Arsenic possesses a similar atomic radius, as well as near identical electronegativity to P (5). The most common form of P in biology is phosphate (PO_4^{3-}), which behaves similarly to arsenate (AsO_4^{3-}) over the range of biologically relevant pH and redox gradients (6). The physicochemical similarity between AsO_4^{3-} and PO_4^{3-} contributes to the biological toxicity of AsO_4^{3-} because metabolic pathways intended for PO_4^{3-} cannot distinguish between the two molecules (7) and AsO_4^{3-} may

be incorporated into some early steps in the pathways [(6) and references therein]. However, it is thought that downstream metabolic processes are generally not compatible with As-incorporating molecules because of differences in the reactivities of P and As compounds (8). These down-

stream biochemical pathways may require the more chemically stable P-based metabolites; the lifetimes of more easily hydrolyzed As-bearing analogs are thought to be too short. However, given the similarities of As and P—and by analogy with trace element substitutions—we hypothesized that AsO_4^{3-} could specifically substitute for PO_4^{3-} in an organism possessing mechanisms to cope with the inherent instability of AsO_4^{3-} compounds (6). Here, we experimentally tested this hypothesis by using AsO_4^{3-} , combined with no added PO_4^{3-} , to select for and isolate a microbe capable of accomplishing this substitution.

Geomicrobiology of GFAJ-1. Mono Lake, located in eastern California, is a hypersaline and alkaline water body with high dissolved arsenic concentrations [200 μM on average (9)]. We used lake sediments as inocula into an aerobic defined artificial medium at pH 9.8 (10, 11) containing 10 mM glucose, vitamins, and trace metals but no added PO_4^{3-} or any additional complex organic supplements (such as yeast extract or peptone), with a regimen of increasing AsO_4^{3-} additions initially spanning the range from 100 μM to 5 mM. These enrichments were taken through many decimal-dilution transfers, greatly reducing any potential carryover of autochthonous phosphorus

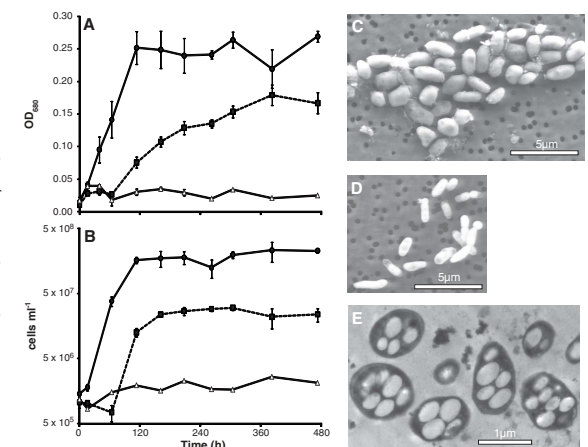


Fig. 1. Growth and electron microscopy of strain GFAJ-1. (A and B) Growth curves of GFAJ-1 grown on the defined synthetic medium amended with either 1.5 mM PO_4^{3-} (solid circles), 40 mM AsO_4^{3-} (solid squares), or neither PO_4^{3-} nor AsO_4^{3-} (open triangles). Cell growth was monitored both by an increase in (A) optical density and (B) cell numbers of the cultures. Symbols represent the mean \pm SD of (A) $n = 6$ experimental and $n = 2$ controls and (B) $n = 3$ experimental and $n = 1$ control. This was a single experiment with six replicates; however, material was conserved to extend the duration of the experiment to allow material for cell-counting samples. (C and D) Scanning electron micrographs of strain GFAJ-1 under two conditions, (C) +As/-P and (D) -As/+P. (E) Transmission electron micrograph of +As/-P GFAJ-1 showed internal vacuole-like structures. Scale bars are as indicated in the figure (11).

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Different Solvent

Metabolism can best operate when metabolites are dissolved

Is water as good a bio-solvent as advertised?

H₂O is a most efficient polar solvent and extremely abundant in the Universe, but also has some drawbacks:

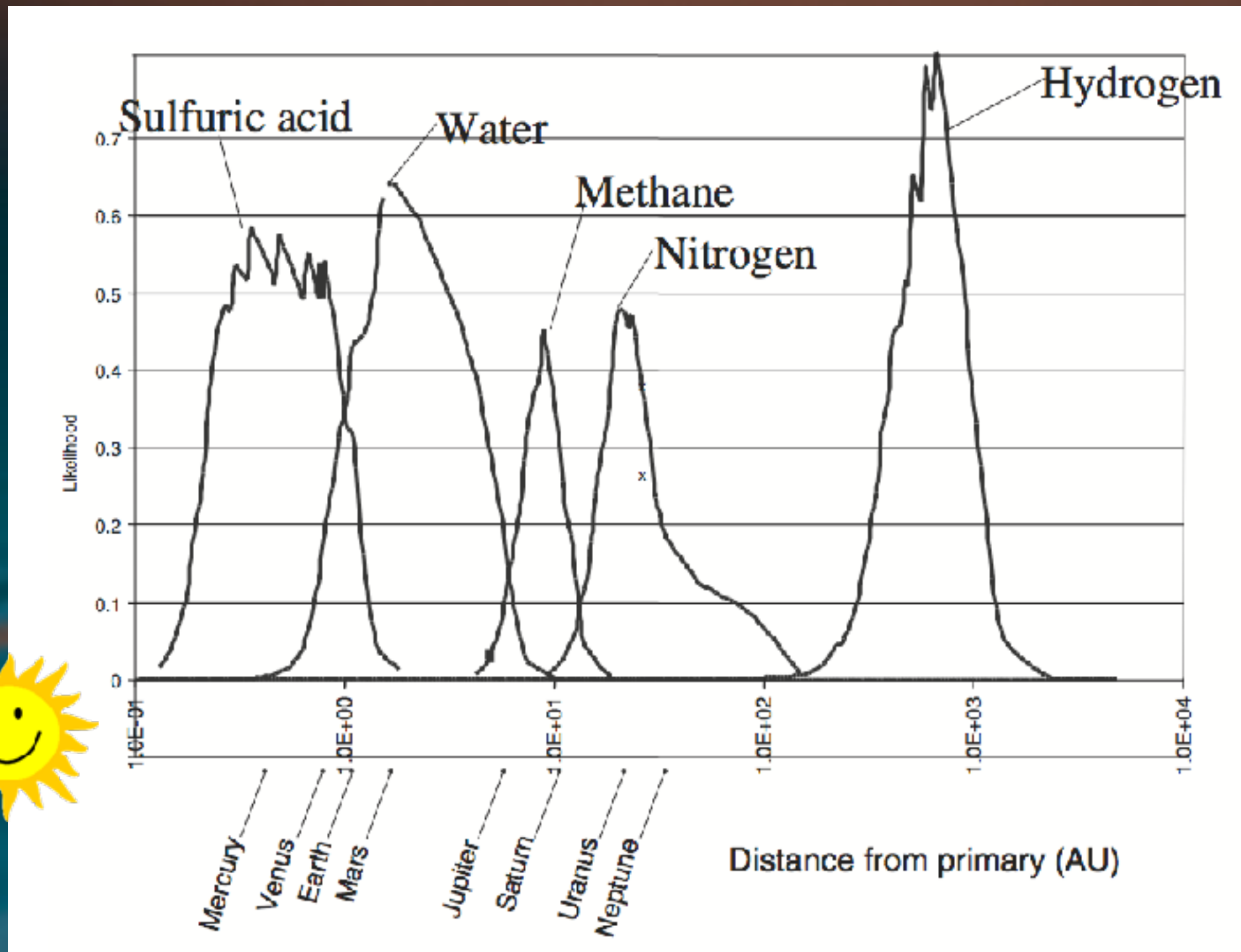
- ❖ It is too good! RNA and DNA are being dissolved in water and need constant repair
- ❖ Liquid water on a planet surface exist only in a narrow distance range from the star (cf. Habitable Zone)

Metabolism can best operate when metabolites are dissolved

The solvent properties (solubility) are used for three tasks beyond dissolving metabolites:

- 1) for compartmentalization (insoluble membrane)
- 2) to build macroscopic structures (e.g. insoluble cellulose)
- 3) to achieve genetic regulation (e.g. selected solubility of steroids)

Alternative solvents



At high temperature, sulfuric acid might work

At low temperature, ammonia (NH_3), nitrogen might work

The alternatives have not extensively been studied in labs yet

Different Elements

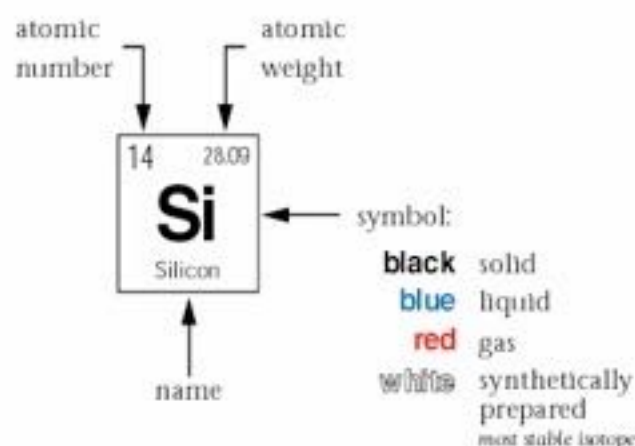
Periodic Table of the Elements



ERNEST ORLANDO LAWRENCE
BERKELEY NATIONAL LABORATORY

Skun3 Seaborg

| | | | | | | | | | | | | | | | | | |
|---|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|--|
| <div><div><div>11.01</div><div>H</div><div>Hydrogen</div></div><div><div>36.94</div><div>Li</div><div>Lithium</div></div><div><div>9.01</div><div>Be</div><div>Beryllium</div></div><div><div>22.99</div><div>Na</div><div>Sodium</div></div><div><div>24.31</div><div>Mg</div><div>Magnesium</div></div><div><div>39.10</div><div>K</div><div>Potassium</div></div><div><div>40.08</div><div>Ca</div><div>Calcium</div></div><div><div>44.96</div><div>Sc</div><div>Scandium</div></div><div><div>47.90</div><div>Ti</div><div>Titanium</div></div><div><div>50.94</div><div>V</div><div>Vanadium</div></div><div><div>51.996</div><div>Cr</div><div>Chromium</div></div><div><div>54.94</div><div>Mn</div><div>Manganese</div></div><div><div>55.85</div><div>Fe</div><div>Iron</div></div><div><div>58.93</div><div>Co</div><div>Cobalt</div></div><div><div>58.70</div><div>Ni</div><div>Nickel</div></div><div><div>63.55</div><div>Cu</div><div>Copper</div></div><div><div>65.37</div><div>Zn</div><div>Zinc</div></div><div><div>69.72</div><div>Ga</div><div>Gallium</div></div><div><div>72.59</div><div>Ge</div><div>Germanium</div></div><div><div>74.92</div><div>As</div><div>Arsenic</div></div><div><div>78.96</div><div>Se</div><div>Selenium</div></div><div><div>79.90</div><div>Br</div><div>Bromine</div></div><div><div>83.80</div><div>Kr</div><div>Krypton</div></div><div><div>85.47</div><div>Rb</div><div>Rubidium</div></div><div><div>87.62</div><div>Sr</div><div>Strontium</div></div><div><div>88.91</div><div>Y</div><div>Yttrium</div></div><div><div>91.22</div><div>Zr</div><div>Zirconium</div></div><div><div>92.91</div><div>Nb</div><div>Niobium</div></div><div><div>95.94</div><div>Mo</div><div>Molybdenum</div></div><div><div>(98)</div><div>Tc</div><div>Technetium</div></div><div><div>101.07</div><div>Ru</div><div>Ruthenium</div></div><div><div>102.91</div><div>Rh</div><div>Rhodium</div></div><div><div>106.40</div><div>Pd</div><div>Palladium</div></div><div><div>107.87</div><div>Ag</div><div>Silver</div></div><div><div>112.41</div><div>Cd</div><div>Cadmium</div></div><div><div>114.82</div><div>In</div><div>Indium</div></div><div><div>118.69</div><div>Sn</div><div>Tin</div></div><div><div>121.75</div><div>Sb</div><div>Antimony</div></div><div><div>127.60</div><div>Te</div><div>Tellurium</div></div><div><div>126.90</div><div>I</div><div>Iodine</div></div><div><div>131.30</div><div>Xe</div><div>Xenon</div></div><div><div>132.91</div><div>Cs</div><div>Cesium</div></div><div><div>137.33</div><div>Ba</div><div>Barium</div></div><div><div>138.91</div><div>La</div><div>Lanthanum</div></div><div><div>178.49</div><div>Hf</div><div>Hafnium</div></div><div><div>180.95</div><div>Ta</div><div>Tantalum</div></div><div><div>183.85</div><div>W</div><div>Tungsten</div></div><div><div>186.21</div><div>Re</div><div>Rhenium</div></div><div><div>190.20</div><div>Os</div><div>Osmium</div></div><div><div>192.22</div><div>Ir</div><div>Iridium</div></div><div><div>196.09</div><div>Pt</div><div>Platinum</div></div><div><div>196.97</div><div>Au</div><div>Gold</div></div><div><div>200.59</div><div>Hg</div><div>Mercury</div></div><div><div>204.37</div><div>Tl</div><div>Thallium</div></div><div><div>207.19</div><div>Pb</div><div>Lead</div></div><div><div>208.98</div><div>Bi</div><div>Bismuth</div></div><div><div>(209)</div><div>Po</div><div>Polonium</div></div><div><div>(210)</div><div>At</div><div>Astatine</div></div><div><div>(222)</div><div>Rn</div><div>Radon</div></div><div><div>(223)</div><div>Fr</div><div>Francium</div></div><div><div>226.03</div><div>Ra</div><div>Radium</div></div><div><div>227.03</div><div>Ac</div><div>Actinium</div></div><div><div>(261)</div><div>Rf</div><div>Rutherfordium</div></div><div><div>(262)</div><div>Ha</div><div>Hahnium</div></div><div><div>(266)</div><div>Sg</div><div>Seaborgium</div></div><div><div>(262)</div><div>Bh</div><div>Bohrium</div></div><div><div>(265)</div><div>Hs</div><div>Hassium</div></div><div><div>(266)</div><div>Mt</div><div>Meitnerium</div></div><div><div>(271)</div><div></div><div></div></div><div><div>(272)</div><div></div><div></div></div><div><div>(277)</div><div></div><div></div></div><div><div>(113)</div><div></div><div></div></div><div><div>(285)</div><div></div><div></div></div><div><div>(115)</div><div></div><div></div></div><div><div>(289)</div><div></div><div></div></div><div><div>(117)</div><div></div><div></div></div><div><div>(293)</div><div></div><div></div></div></div> <div><div>21.01</div><div>H</div><div>Hydrogen</div></div> <div><div>4.003</div><div>He</div><div>Helium</div></div> <div><div>20.18</div><div>Ne</div><div>Neon</div></div> <div><div>39.95</div><div>Ar</div><div>Argon</div></div> <div><div>83.80</div><div>Kr</div><div>Krypton</div></div> <div><div>131.30</div><div>Xe</div><div>Xenon</div></div> <div><div>222</div><div>Rn</div><div>Radon</div></div> <div><div>14</div><div>28.09</div><div>Si</div><div>Silicon</div></div> <div><div>atomic number</div><div>atomic weight</div><div>symbol:</div><div>black solid</div><div>blue liquid</div><div>red gas</div><div>white synthetically prepared most stable isotope</div><div>name</div></div> <div><div>alkali metals</div><div>alkaline earth metals</div><div>transitional metals</div><div>other metals</div><div>nonmetals</div><div>noble gases</div></div> <div><div>5</div><div>10.81</div><div>B</div><div>Boron</div></div> <div><div>6</div><div>12.01</div><div>C</div><div>Carbon</div></div> <div><div>7</div><div>14.01</div><div>N</div><div>Nitrogen</div></div> <div><div>8</div><div>15.999</div><div>O</div><div>Oxygen</div></div> <div><div>9</div><div>18.998</div><div>F</div><div>Fluorine</div></div> <div><div>10</div><div>20.18</div><div>Ne</div><div>Neon</div></div> <div><div>11</div><div>22.99</div><div>Na</div><div>Sodium</div></div> <div><div>12</div><div>24.31</div><div>Mg</div><div>Magnesium</div></div> <div><div>13</div><div>26.98</div><div>Al</div><div>Aluminum</div></div> <div><div>14</div><div>28.09</div><div>Si</div><div>Silicon</div></div> <div><div>15</div><div>30.97</div><div>P</div><div>Phosphorus</div></div> <div><div>16</div><div>32.06</div><div>S</div><div>Sulfur</div></div> <div><div>17</div><div>35.45</div><div>Cl</div><div>Chlorine</div></div> <div><div>18</div><div>39.95</div><div>Ar</div><div>Argon</div></div> <div><div>19</div><div>39.10</div><div>K</div><div>Potassium</div></div> <div><div>20</div><div>40.08</div><div>Ca</div><div>Calcium</div></div> <div><div>21</div><div>44.96</div><div>Sc</div><div>Scandium</div></div> <div><div>22</div><div>47.90</div><div>Ti</div><div>Titanium</div></div> <div><div>23</div><div>50.94</div><div>V</div><div>Vanadium</div></div> <div><div>24</div><div>51.996</div><div>Cr</div><div>Chromium</div></div> <div><div>25</div><div>54.94</div><div>Mn</div><div>Manganese</div></div> <div><div>26</div><div>55.85</div><div>Fe</div><div>Iron</div></div> <div><div>27</div><div>58.93</div><div>Co</div><div>Cobalt</div></div> <div><div>28</div><div>58.70</div><div>Ni</div><div>Nickel</div></div> <div><div>29</div><div>63.55</div><div>Cu</div><div>Copper</div></div> <div><div>30</div><div>65.37</div><div>Zn</div><div>Zinc</div></div> <div><div>31</div><div>69.72</div><div>Ga</div><div>Gallium</div></div> <div><div>32</div><div>72.59</div><div>Ge</div><div>Germanium</div></div> <div><div>33</div><div>74.92</div><div>As</div><div>Arsenic</div></div> <div><div>34</div><div>78.96</div><div>Se</div><div>Selenium</div></div> <div><div>35</div><div>79.90</div><div>Br</div><div>Bromine</div></div> <div><div>36</div><div>83.80</div><div>Kr</div><div>Krypton</div></div> <div><div>37</div><div>85.47</div><div>Rb</div><div>Rubidium</div></div> <div><div>38</div><div>87.62</div><div>Sr</div><div>Strontium</div></div> <div><div>39</div><div>88.91</div><div>Y</div><div>Yttrium</div></div> <div><div>40</div><div>91.22</div><div>Zr</div><div>Zirconium</div></div> <div><div>41</div><div>92.91</div><div>Nb</div><div>Niobium</div></div> <div><div>42</div><div>95.94</div><div>Mo</div><div>Molybdenum</div></div> <div><div>43</div><div>(98)</div><div>Tc</div><div>Technetium</div></div> <div><div>44</div><div>101.07</div><div>Ru</div><div>Ruthenium</div></div> <div><div>45</div><div>102.91</div><div>Rh</div><div>Rhodium</div></div> <div><div>46</div><div>106.40</div><div>Pd</div><div>Palladium</div></div> <div><div>47</div><div>107.87</div><div>Ag</div><div>Silver</div></div> <div><div>48</div><div>112.41</div><div>Cd</div><div>Cadmium</div></div> <div><div>49</div><div>114.82</div><div>In</div><div>Indium</div></div> <div><div>50</div><div>118.69</div><div>Sn</div><div>Tin</div></div> <div><div>51</div><div>121.75</div><div>Sb</div><div>Antimony</div></div> <div><div>52</div><div>127.60</div><div>Te</div><div>Tellurium</div></div> <div><div>53</div><div>126.90</div><div>I</div><div>Iodine</div></div> <div><div>54</div><div>131.30</div><div>Xe</div><div>Xenon</div></div> <div><div>55</div><div>132.91</div><div>Cs</div><div>Cesium</div></div> <div><div>56</div><div>137.33</div><div>Ba</div><div>Barium</div></div> <div><div>57</div><div>138.91</div><div>La</div><div>Lanthanum</div></div> <div><div>72</div><div>178.49</div><div>Hf</div><div>Hafnium</div></div> <div><div>73</div><div>180.95</div><div>Ta</div><div>Tantalum</div></div> <div><div>74</div><div>183.85</div><div>W</div><div>Tungsten</div></div> <div><div>75</div><div>186.21</div><div>Re</div><div>Rhenium</div></div> <div><div>76</div><div>190.20</div><div>Os</div><div>Osmium</div></div> <div><div>77</div><div>192.22</div><div>Ir</div><div>Iridium</div></div> <div><div>78</div><div>196.09</div><div>Pt</div><div>Platinum</div></div> <div><div>79</div><div>196.97</div><div>Au</div><div>Gold</div></div> <div><div>80</div><div>200.59</div><div>Hg</div><div>Mercury</div></div> <div><div>81</div><div>204.37</div><div>Tl</div><div>Thallium</div></div> <div><div>82</div><div>207.19</div><div>Pb</div><div>Lead</div></div> <div><div>83</div><div>208.98</div><div>Bi</div><div>Bismuth</div></div> <div><div>84</div><div>(209)</div><div>Po</div><div>Polonium</div></div> <div><div>85</div><div>(210)</div><div>At</div><div>Astatine</div></div> <div><div>86</div><div>(222)</div><div>Rn</div><div>Radon</div></div> <div><div>87</div><div>(223)</div><div>Fr</div><div>Francium</div></div> <div><div>88</div><div>226.03</div><div>Ra</div><div>Radium</div></div> <div><div>89</div><div>227.03</div><div>Ac</div><div>Actinium</div></div> <div><div>104</div><div>(261)</div><div>Rf</div><div>Rutherfordium</div></div> <div><div>105</div><div>(262)</div><div>Ha</div><div>Hahnium</div></div> <div><div>106</div><div>(266)</div><div>Sg</div><div>Seaborgium</div></div> <div><div>107</div><div>(262)</div><div>Bh</div><div>Bohrium</div></div> <div><div>108</div><div>(265)</div><div>Hs</div><div>Hassium</div></div> <div><div>109</div><div>(266)</div><div>Mt</div><div>Meitnerium</div></div> <div><div>110</div><div>(271)</div><div></div><div></div></div> <div><div>111</div><div>(272)</div><div></div><div></div></div> <div><div>112</div><div>(277)</div><div></div><div></div></div> <div><div>(113)</div><div></div><div></div></div> <div><div>(285)</div><div></div><div></div></div> <div><div>(115)</div><div></div><div></div></div> <div><div>(289)</div><div></div><div></div></div> <div><div>(117)</div><div></div><div></div></div> <div><div>(293)</div><div></div><div></div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> 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<div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> <div><div>Si</div><div>Silicon</div></div> 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- alkali metals
- alkaline earth metals
- transitional metals
- other metals
- nonmetals
- noble gases

Lanthanide series ▶

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| 58 140.12 Ce Cerium | 59 140.91 Pr Praseodymium | 60 144.24 Nd Neodymium | 61 (145) Pm Promethium | 62 150.40 Sm Samarium | 63 151.96 Eu Europium | 64 157.25 Gd Gadolinium | 65 158.93 Tb Terbium | 66 162.50 Dy Dysprosium | 67 164.93 Ho Holmium | 68 167.26 Er Erbium | 69 168.93 Tm Thulium | 70 173.04 Yb Ytterbium | 71 174.97 Lu Lutetium |
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Actinide series ▶

| | | | | | | | | | | | | | |
|--------------------------------------|---|-------------------------------------|--|---------------------------------------|---------------------------------------|------------------------------------|---------------------------------------|---|---|--------------------------------------|--|---------------------------------------|---|
| 90 232.04 Th Thorium | 91 231.04 Pa Protactinium | 92 238.03 U Uranium | 93 237.05 Np Neptunium | 94 (244) Pu Plutonium | 95 (243) Am Americium | 96 (247) Cm Curium | 97 (247) Bk Berkelium | 98 (251) Cf Californium | 99 (252) Es Einsteinium | 100 (257) Fm Fermium | 101 (260) Md Mendelevium | 102 (259) No Nobelium | 103 (262) Lr Lawrencium |
|--------------------------------------|---|-------------------------------------|--|---------------------------------------|---------------------------------------|------------------------------------|---------------------------------------|---|---|--------------------------------------|--|---------------------------------------|---|

This is getting weird...

Terran life is build on CHONPS

Ammonia as solvent would require N=C bonds (while water as solvent builds on O=C bonds): CHONPS → CHNOPPS

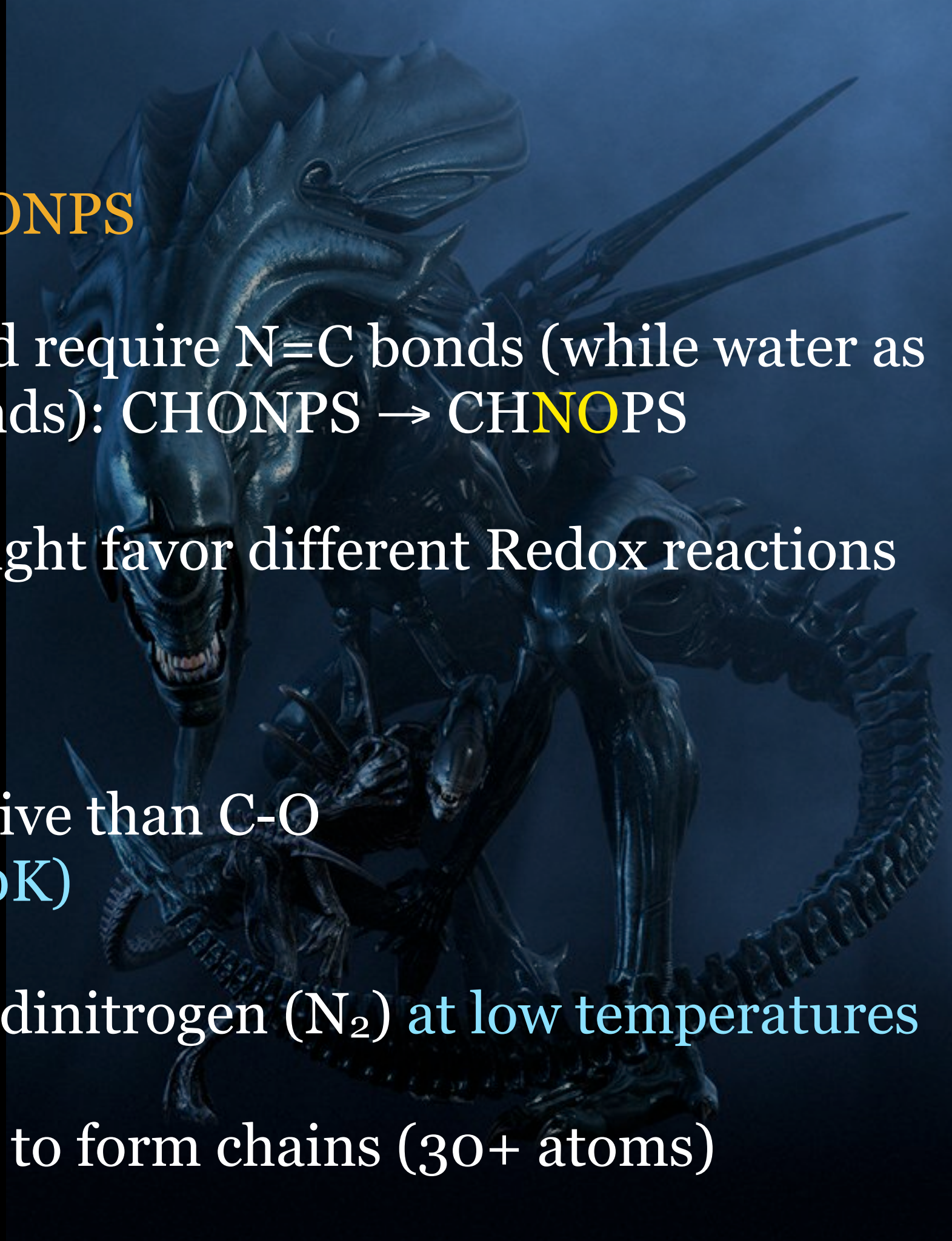
A different atmosphere might favor different Redox reactions

Silicon life:

Si-O bonds are more reactive than C-O
at low temperatures (<270K)

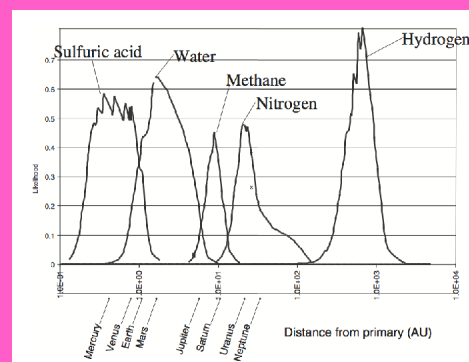
Si-O bonds are soluble in dinitrogen (N₂) at low temperatures

Si-Si double bonds known to form chains (30+ atoms)



M Dwarfs are the most common stars in the Milky Way

How far does their “N₂ Habitable Zone” extend?



Different Architecture

Life “Classic”
Life “RNA World”
Life “Weird”

Life “Classic”

Multiple biopolymers:

- ♦ DNA used for genetics
- ♦ Proteins used to solve structural and catalytic problems
- ♦ RNA is encoded to act as intermediate between the two

Life “RNA World”

Single biopolymer:

- ♦ Ribozymes does it all

A single biopolymer is a lot simpler - this type of life might be a lot more abundant in the Universe...

Life “Classic”
Life “RNA World”
Life “Weird”

Life “Weird”

- ✦ Different single biopolymer perhaps driven by lack of space or scarce resources...

Arguments for at least two biopolymers would be:

- ✦ The catalyst should be optimized for many chemical reactions, i.e. easily change its physical properties
- ✦ The genetic molecule should be as simple as possible to allow accurate replication. It should be stable to allow Darwinian evolution

Very different Life

What would life look like on Jupiter?

Sagan & Salpeter 1976 imagined what it would take to have living beings in Jupiter's atmosphere: **Sinkers and Floaters**

The challenge: survive in the ammonia troposphere (below the H/He thermo- and stratosphere)

The idea:

Sinkers - tiny hydrogen filled organisms, floating in the upper troposphere

Coalescence makes them to **Floaters** - lighter than environment, while evolution would add sense organs and directional flight, eventually making them **hunters**

Could machines take over life?

Computers take over more and more tasks in our lives..

When will computer be so fast and complex, that we cannot differentiate them from being conscious?

Will computers program themselves to specialize, collect resources and build new, more complex generations?



What is the current “speed” of
the brain compared to
computers?

2010

Internet
1 quintillion bytes

●
Data storage

🌀
Processing speed
(megaflops = million
operations per second)

💡
Power consumption
(1 LED flashlight
bulb = 1 watt)



Cat Brain
98 trillion bytes
61 million megaflops



iPad 2
64 billion bytes
170 megaflops
2.5 watts



**Human
Genome**
750 million bytes

Fastest Supercomputer

(K computer, Fujitsu)

30 quadrillion bytes
8.2 billion megaflops
9.9 million watts

~ 10^{16}
2023: 10^{18}

9.9 million LEDs (not all shown)

Human Brain

3.5 quadrillion bytes
2.2 billion megaflops
20 watts

Life beyond our Universe

Multiverses are a possibility



Life would develop under different laws of physics

Deviations from the standard model of particles physics

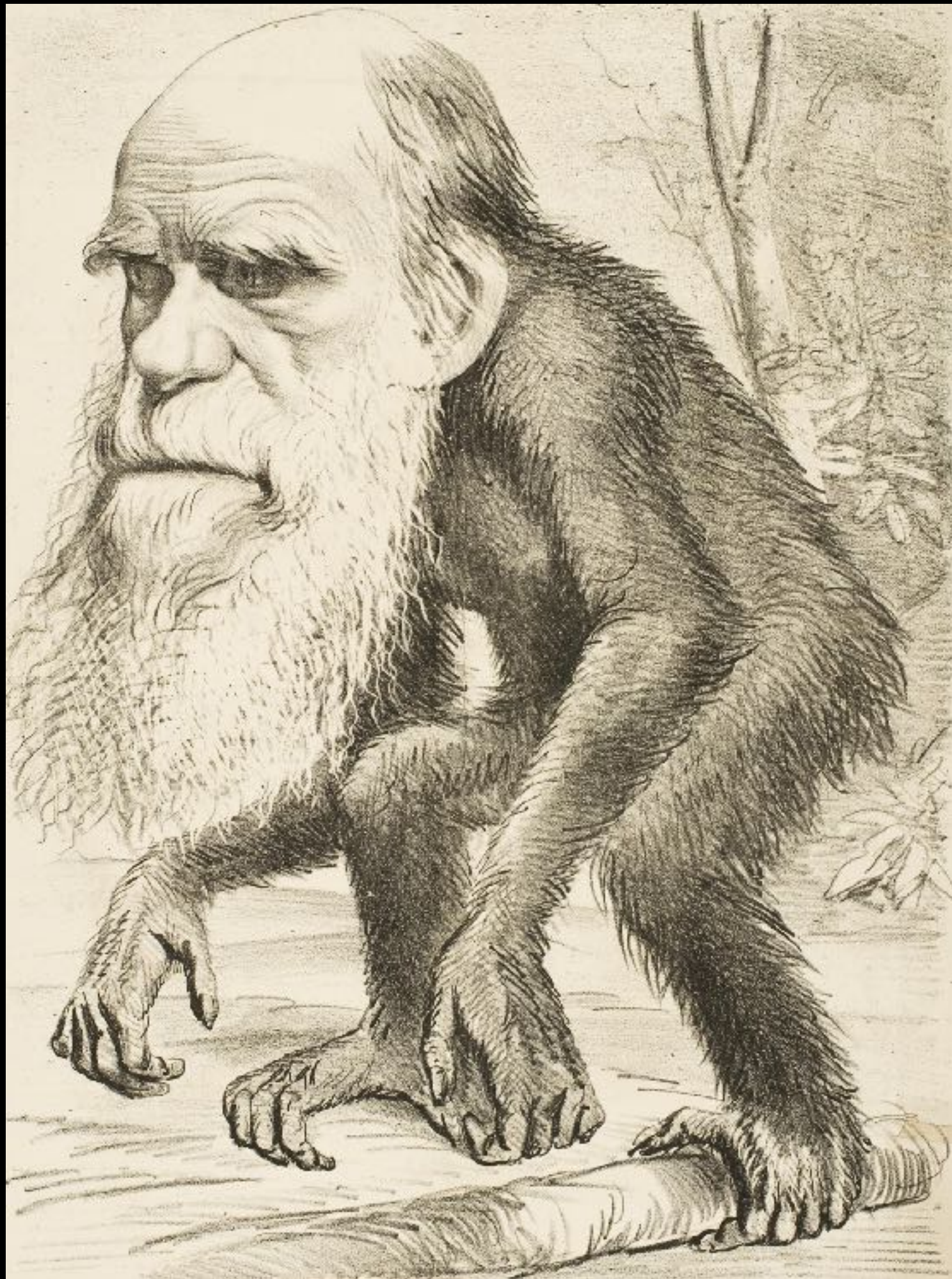
Changing Gravity: changes cosmology, structure formation, evolution of the Universe as a whole...

Changing the masses of Quarks: changes atomic physics and chemistry, star formation... does not exclude “life”

Removing the Weak Force: Big Bang nucleosynthesis would be overabundant in H and He - only massive, short lived stars... does not exclude “life”

More/less dimensions: e.g. *Flatland* by E. Abbott (1884)

The Evolution of Intelligence



It took us a while to get there...

Domain: Eukarya (nucleated cells, aerobic metabolism) [~ 2 Ga ago]

Kingdom: Animalia (locomotion, oxygen respiration, sexual reproduction, large number of different cells)

Phylum: Chordata (body plan, centralized nervous system)

Class: Mammalia (hair, breast feeding) [~ 65 Ma ago]

Order: Primates (dexterity, acute vision, large brains)

Family: Hominidae (complex behaviour, large body size)

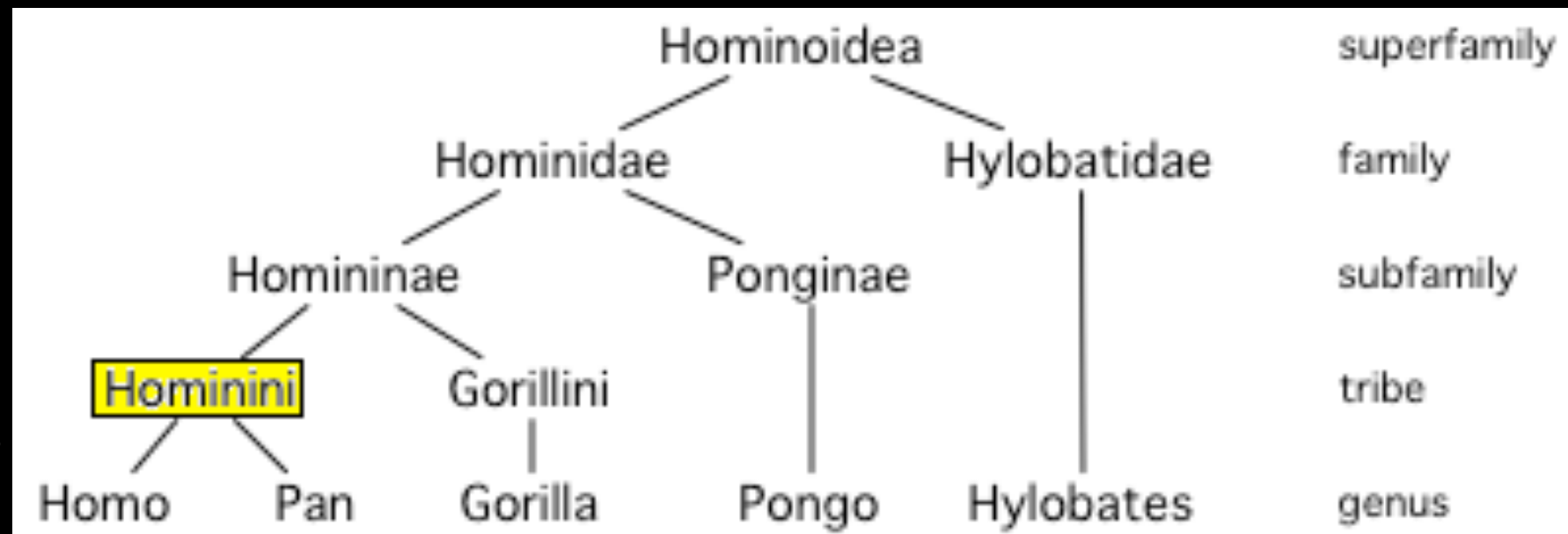
Genus: Homo [~ 2.5 Ma ago]

Species: Sapiens [~ 300.000 a ago]



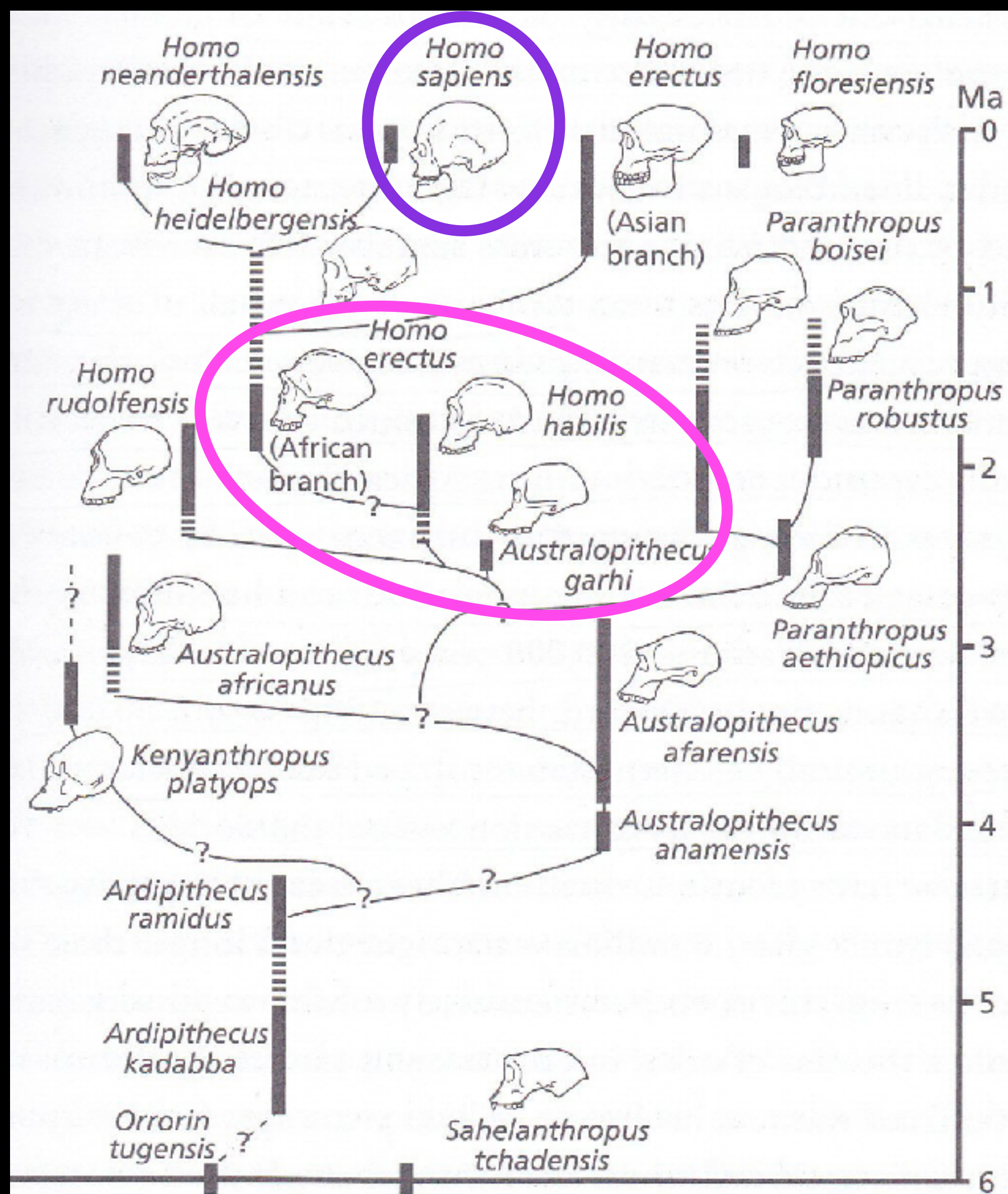
~14 Ma ago

~6 Ma ago

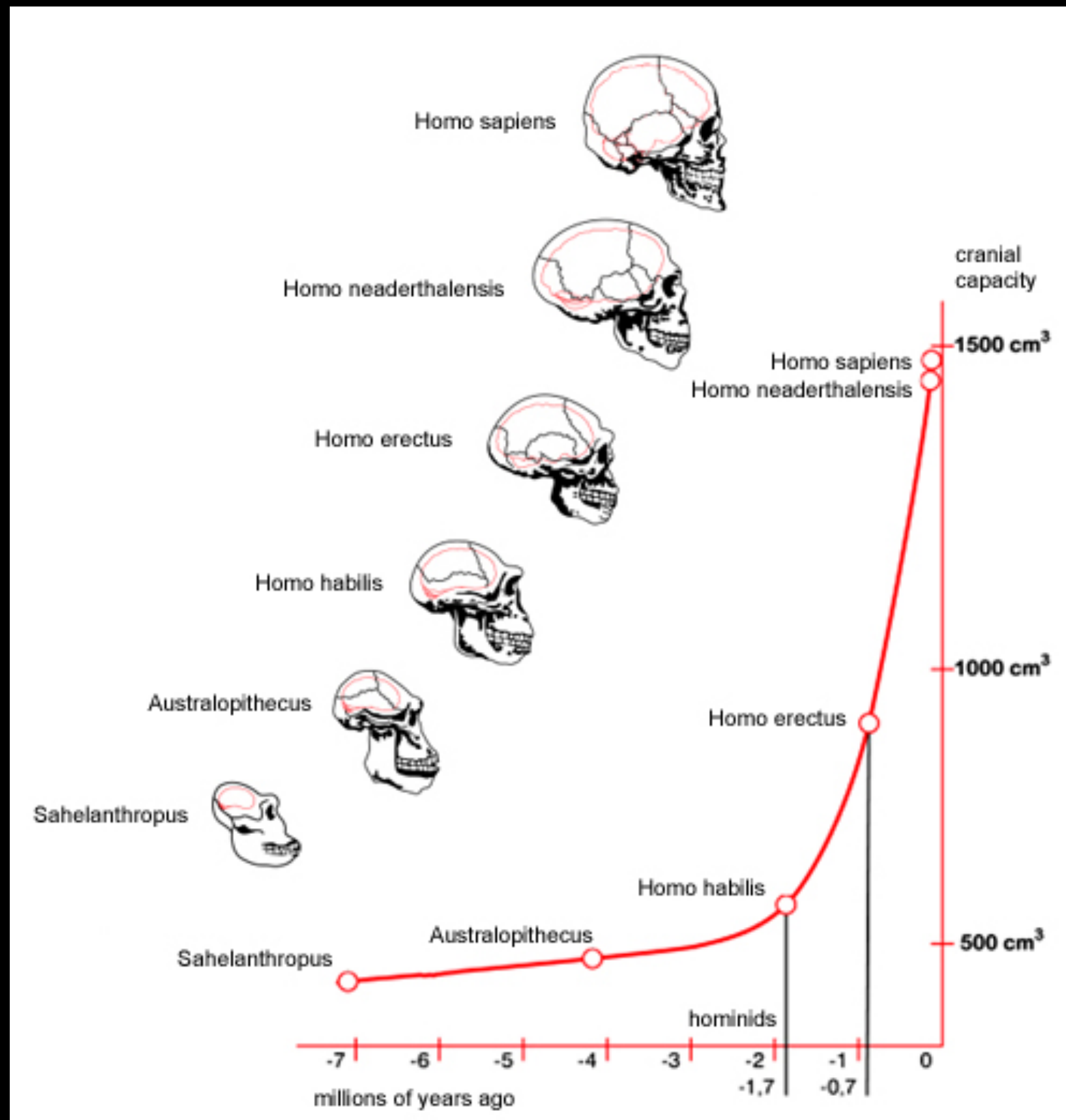


Homo Sapiens appeared
~300.000 years ago
(until recently believed
~150.000 years ago)

mtDNA pointed to
single origin (East
Africa); but oldest now
found in Morocco...



The **brain size** evolved incredibly rapidly after the Australopithecus (in the last 1-2 million years)



A series of pre-adaptations (each could have gone wrong...):

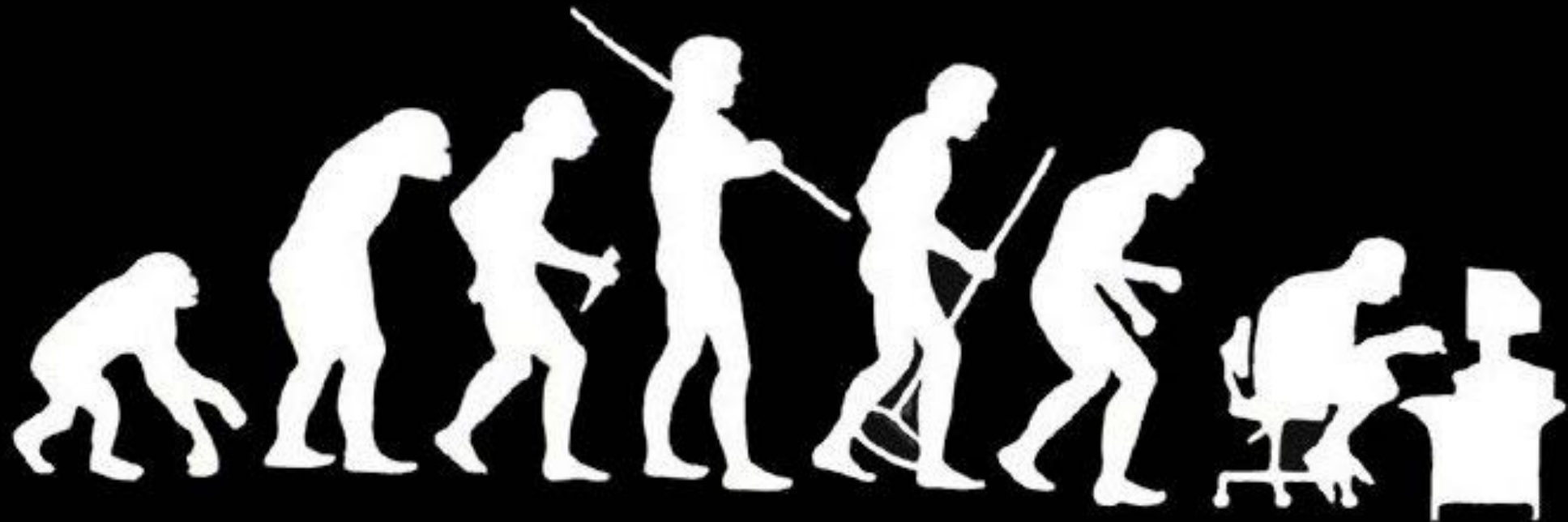
- **Live on land** - enables later use of tools and fire
- **Body size** - enable complex brain (of >1kg)
- **Grasping hands** (no claws) - manipulate objects, further develop the brain, not use for locomotion
- **Mix/Meat diet** - more energy, more proteins to develop the brain; hunting: group activity
[homo erectus - 1 million years ago]
- **Controlled use of fire** - cooked meat, better tool; cooking as social activity promoted groups/camps/“nest”; self-organization of the group

In such eusocial groups, social intelligence is a strong evolutionary advantage



The Evolution of the brain is most likely due to a step by step evolution of social behavior for survival and pressure on social intelligence.

See E.O.Wilson: “The Social Conquest of Earth”



Something, somewhere went terribly wrong

NEW YORK TIMES BESTSELLER

THE SOCIAL CONQUEST OF EARTH



EDWARD
O. WILSON

WINNER OF THE PULITZER PRIZE

"A sweeping argument about the biological origins of complex human culture. . . .
Well-crafted and captivating." —Michael Cavanaugh, *Wall Street Journal*

"*Sapiens* tackles the biggest questions of history and of the modern world, and it is written in unforgettably vivid language."

—JARED DIAMOND, Pulitzer Prize-winning
author of *Guns, Germs, and Steel*

Yuval Noah Harari

Sapiens

A Brief
History of
Humankind

THE PULITZER PRIZE WINNER

CARL SAGAN

Coauthor of *SHADOWS OF FORGOTTEN ANCESTORS*

THE DRAGONS OF EDEN



SPECULATIONS ON THE EVOLUTION
OF HUMAN INTELLIGENCE

"WILL LEAVE THE READER EXHILARATED AND TINGLING...A MASTERPIECE"

—*Chicago Tribune Book World*

Take a break...

Intelligence (planning complex actions, sophisticated operations, elaborate communication) is widespread and easy to explain by evolutionary pressure

Consciousness / Self-awareness (“theatre of the mind”) allows to consider alternatives and run multiple times situations in abstract form - eventually enabling even more complex behaviour

Astrobiology: consciousness allows to contemplate the Universe and to search for life elsewhere

Does consciousness always accompany intelligence?



How many intelligent species
live on Earth?

Human intelligence forms:

<https://www.bundesdruckerei.de/de/innovation-hub/was-ist-intelligenz>

Climate changes - future perspectives

The emergence of civilisation (including leisure time for art and sciences) is related to the stability of the climate after the last glaciation (~15.000 to ~11.000 years ago)

The **agricultural lifestyle** (starting ~8.000 ago) was key to this development

Partly as a result, 8 millennia later, our species started to interact with its environment/climate

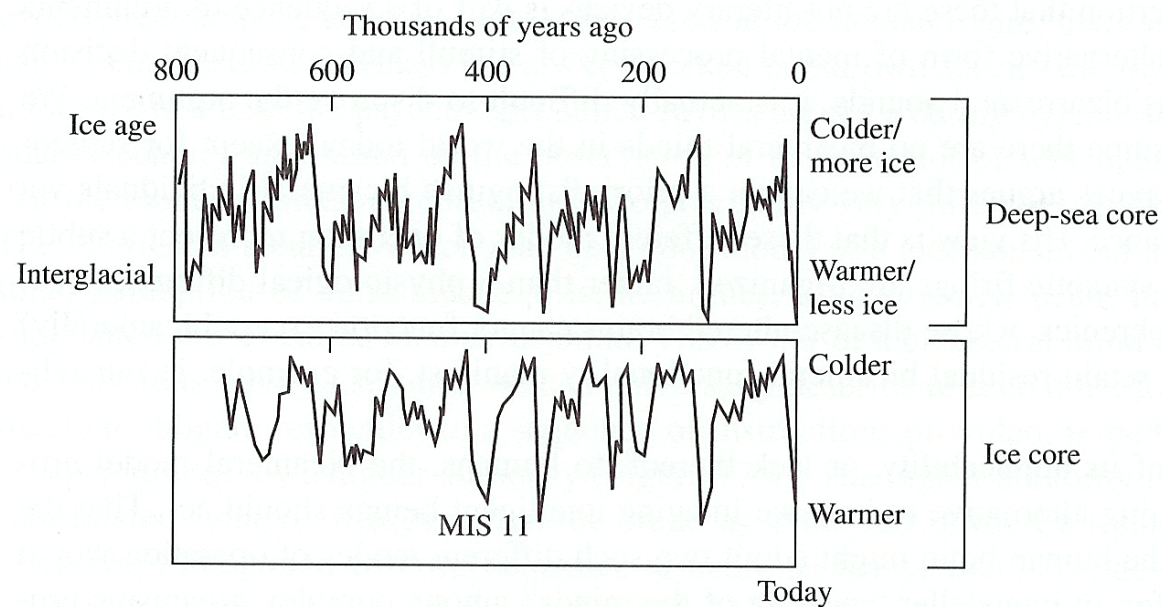


FIGURE 17.5 Pleistocene climate derived from the abundance of deuterium versus hydrogen in an Antarctic ice core, compared with the climate derived from oxygen isotopes in deep sea sediments. The record goes back 800,000 years, and shows among other features an interglacial at 400,000 years ("MIS-11") that lasted longer than the Holocene.



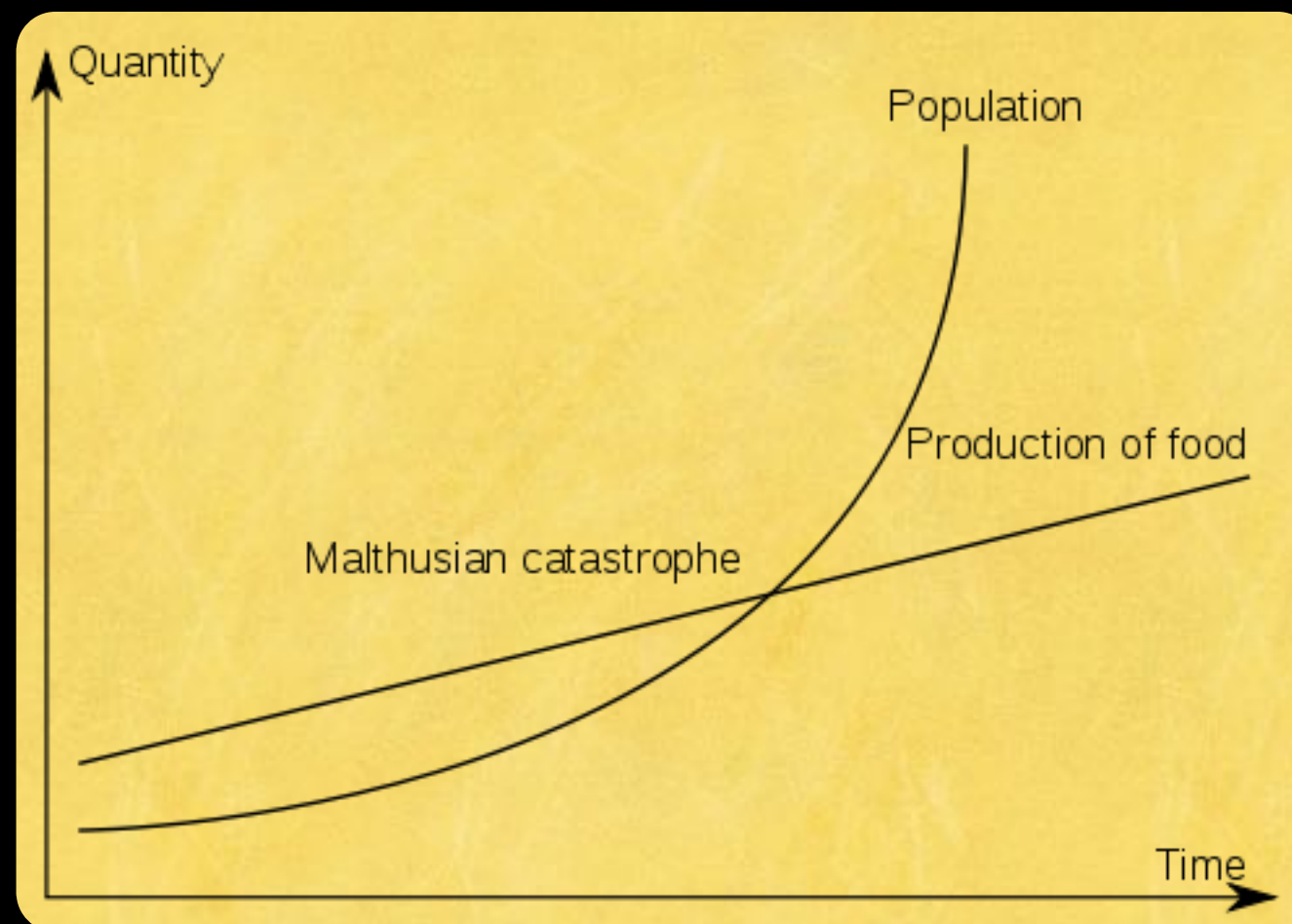
The typical lifetime of an animal species is 1 to 10 million years

The Homo exists since ~300.000 years...

Agricultures and cities since ~10.000 years...

The Challenge: make the last three centuries a prelude to a long-lasting civilization rather than to a Malthusian catastrophe

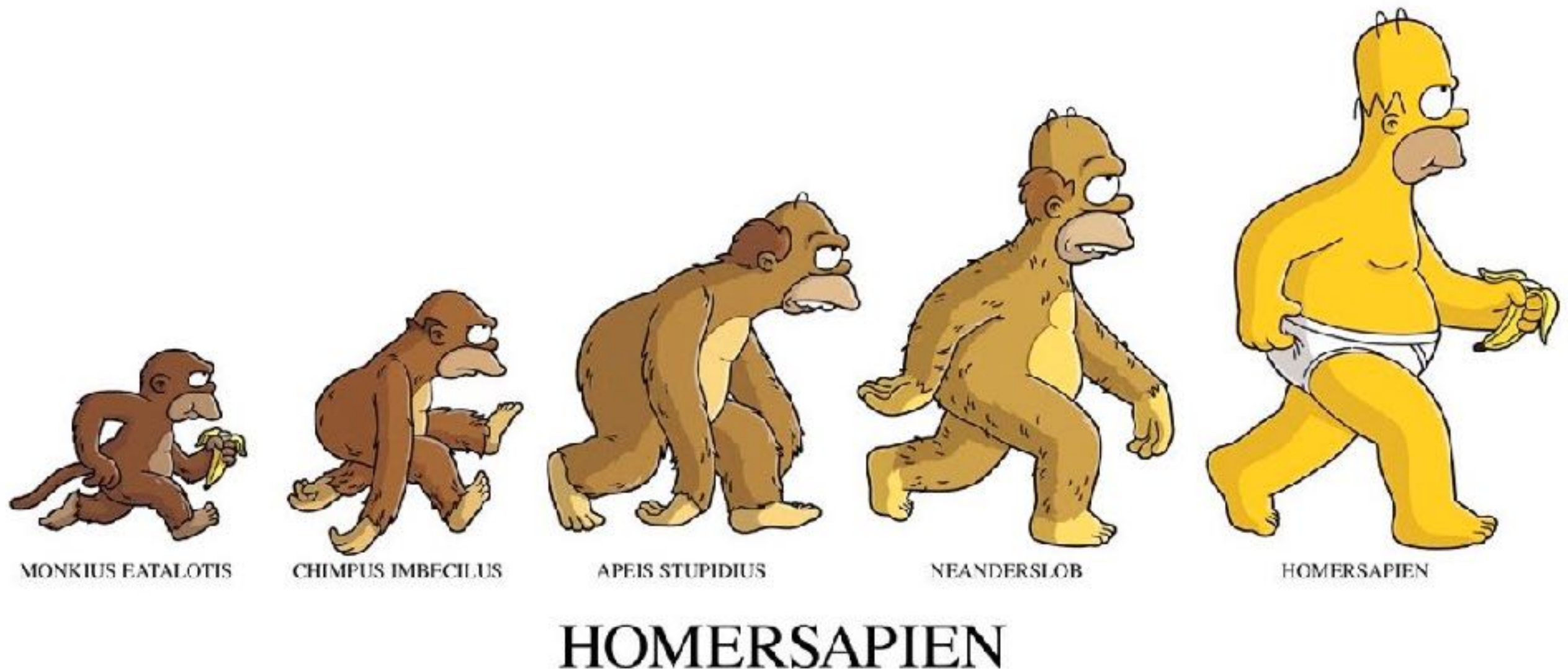
Sagan: we might become space farers not by choice but by necessity



And Humans?

We already think of ourselves as “post-Darwinian” organisms

What will happen once we start genome engineering?



What defined “Darwinian”
evolution?

Will Humans still follow
“Darwinian” evolution?

naturevideo

visualization by CultSci

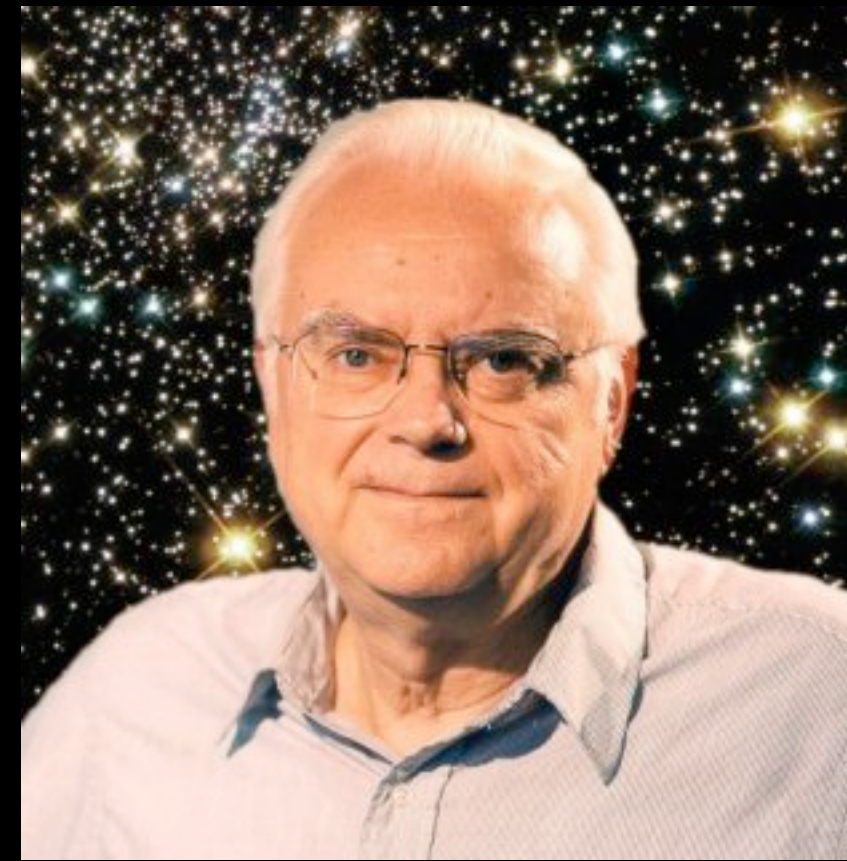
The Drake equation

The Drake Equation

1961, in preparation of the first meeting on detecting extraterrestrial intelligence...

SETI became a research field

N: the number of civilizations in our galaxy with which communication might be possible



Frank Drake (1930-)

$$N = \text{SFR} \times f_{\text{planets}} \times f_{\text{habitable}} \times f_{\text{habited}} \times f_{\text{intel}} \times f_{\text{tech}} \times L$$

The Drake Equation

$$N = \text{SFR} \times f_{\text{planets}} \times n_{\text{habitable}} \times f_{\text{habited}} \times f_{\text{intel}} \times f_{\text{tech}} \times L$$

Original estimates:

SFR=10/year, f_p =50%, n_h =2, f_h =100%, f_i =1%, f_t =1%, $L=10^4$ years

$N = 10$

From this lecture, with optimism:

SFR=10/year, f_p =50%, n_h =2, f_h =50%, f_i =100%, f_t =100%, $L=10^4$ years

$N = 50.000$ (statistically the closest is <100 pc away)

but N could be anywhere between 0 and 10^{6+}

Compute your favorite N from the Drake equation

$$N = \text{SFR} \times f_{\text{planets}} \times n_{\text{habitable}} \times f_{\text{habited}} \times f_{\text{intel}} \times f_{\text{tech}} \times L$$

⇒ *Fermi paradox*

but sufficient drive for SETI



Searching for Extraterrestrial Intelligence

SETI

Search for Extra-Terrestrial Intelligence

Intelligence is not well defined on Earth...

What are we searching for on other planets?



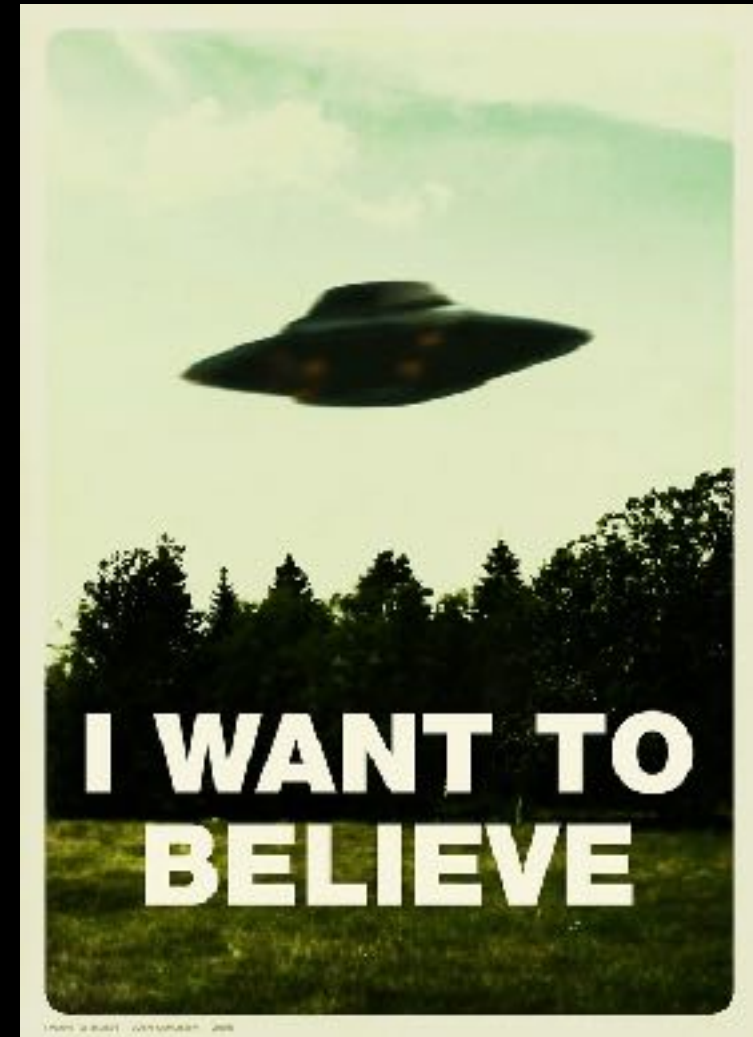
“Intelligence” = Technology?!

We assume that a civilization will develop:

- Energy production
- Transportation
- Information exchange
- (and Waging war)

Scenario 1: the civilisation is **local**
→ they do not try to communicate
⇒ look for unintentional manifestations

Scenario 2: they intend interstellar **information**
exchange (travel?)



Fermi paradox (“Where is everybody?”)

Solutions:

- a)** they are here (e.g. directed panspermia)
- b)** they exist but have not yet communicated
(or we failed to understand the signal)
- c)** they do not exist

In all case, rules for **efficient communication** seem to be:

- ❖ have the maximum possible velocity
- ❖ be easy to generate, launch and capture
- ❖ not be absorbed by the interstellar medium
- ❖ require minimum energy per bit of information (?)
- ❖ go where aimed (?)

Consequently, communication focuses on **electromagnetic radiation**



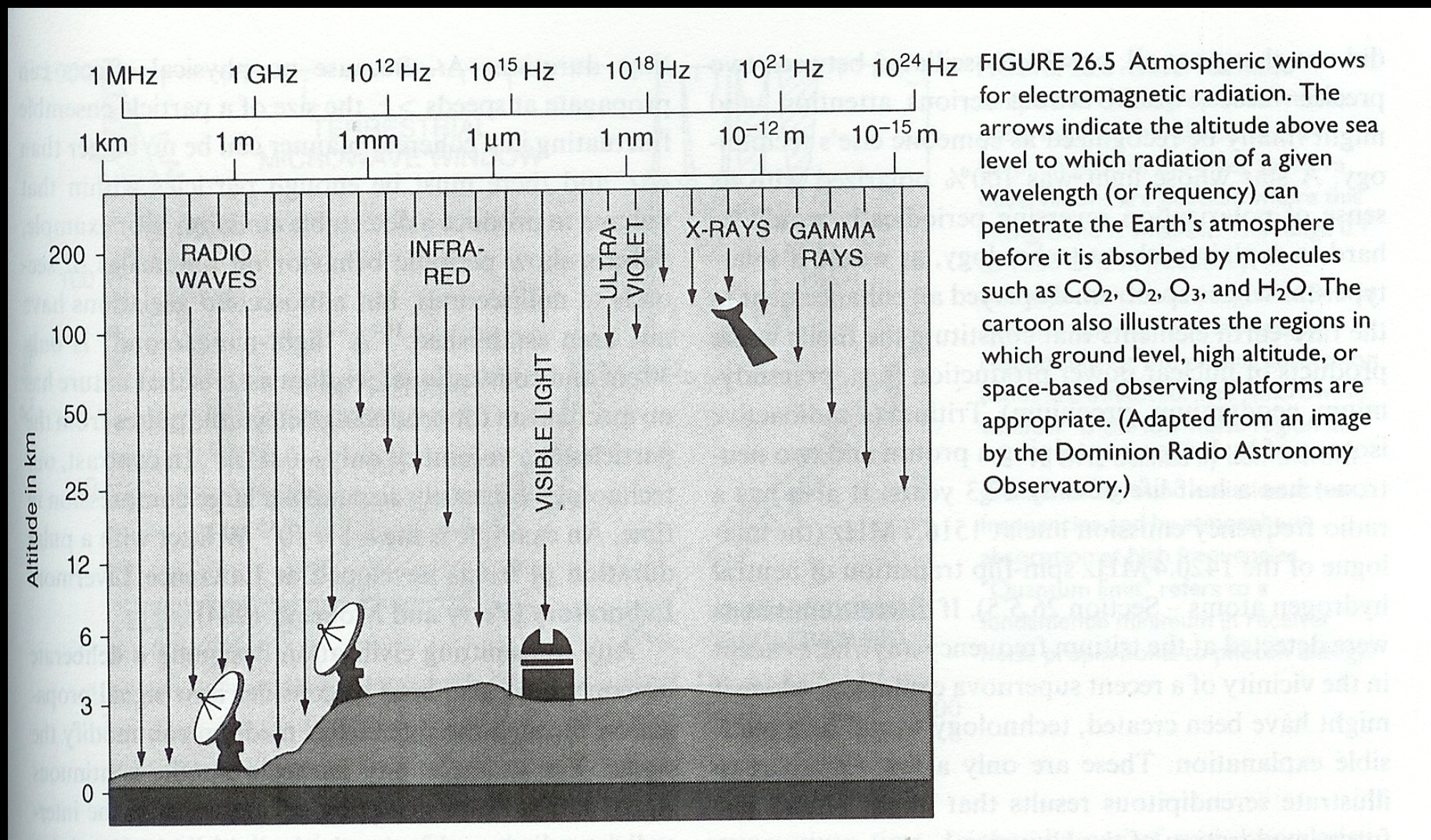
Search Strategies

At which frequencies do we have a chance of detecting signal?

Astrophysical background: the sky is bright from the radio to γ -rays

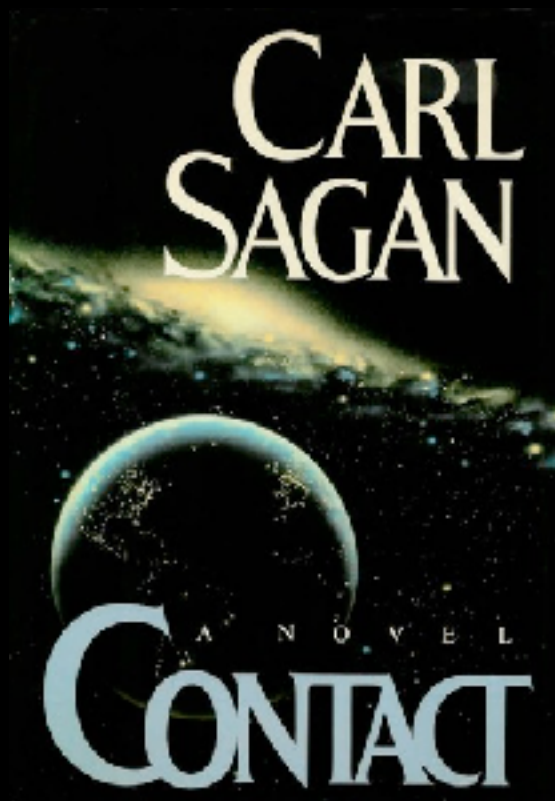
In practice: spatial, spectral and temporal filters are used to reduce that background

The Earth atmosphere limits the accessible wavelength range



Can we tell natural from artificial signals?

- a) the signal could be unrecognisable for our (not very advanced) technology [tough luck...]
- b) the signal would mimic an astrophysical signal (e.g. modulated pulsar, 100% polarised star, ...)
- c) the signal cannot be produced by an astrophysical source (e.g. small $\Delta\nu$, or small Δt) [most SETI searches]

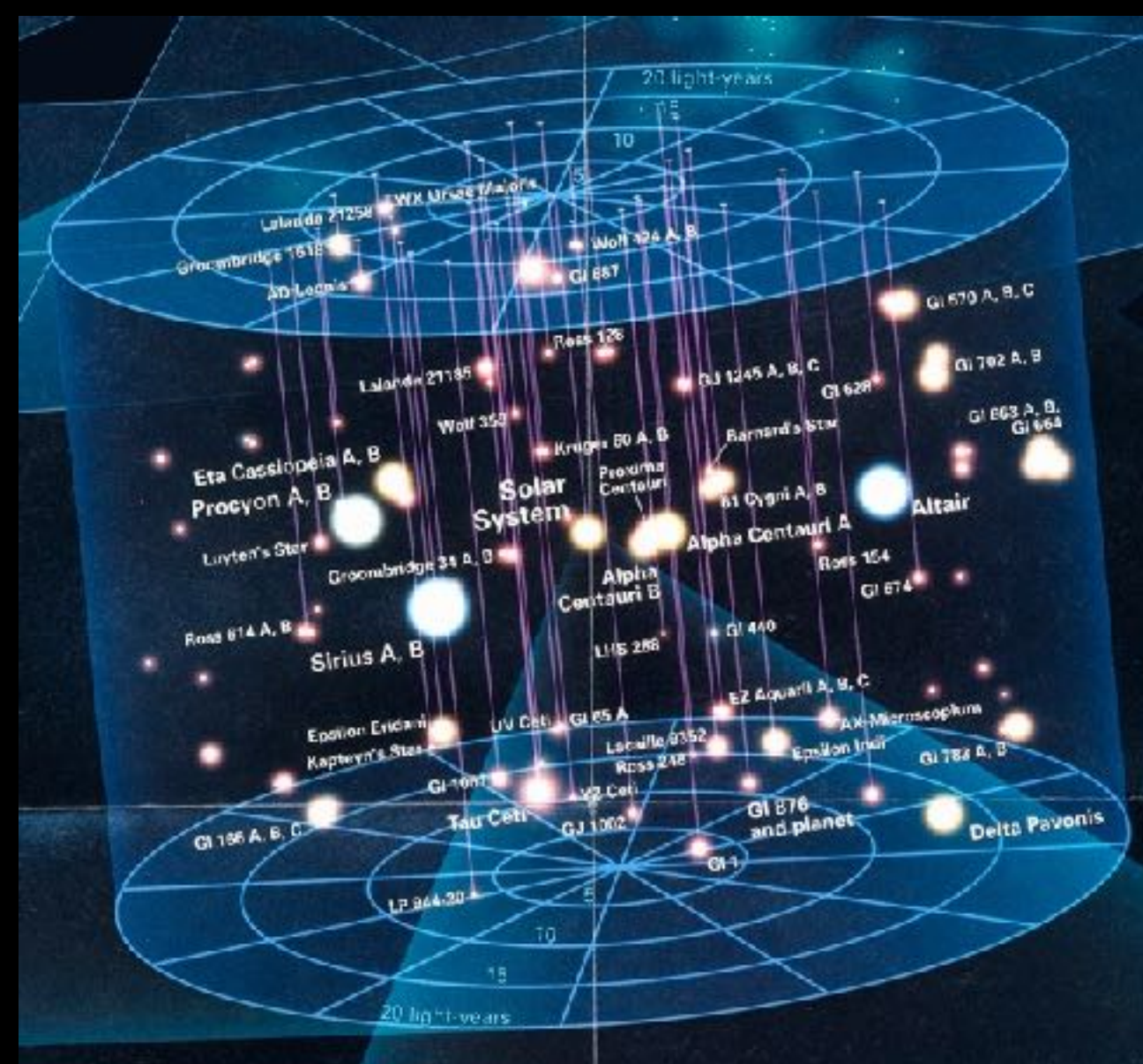


Targets or Sky sweeps?

Focused strategies usually target G2V stars (sun analogs)

Sky surveys suffer from lower sensitivity and irregular coverage

No strategy was successful so far 😊

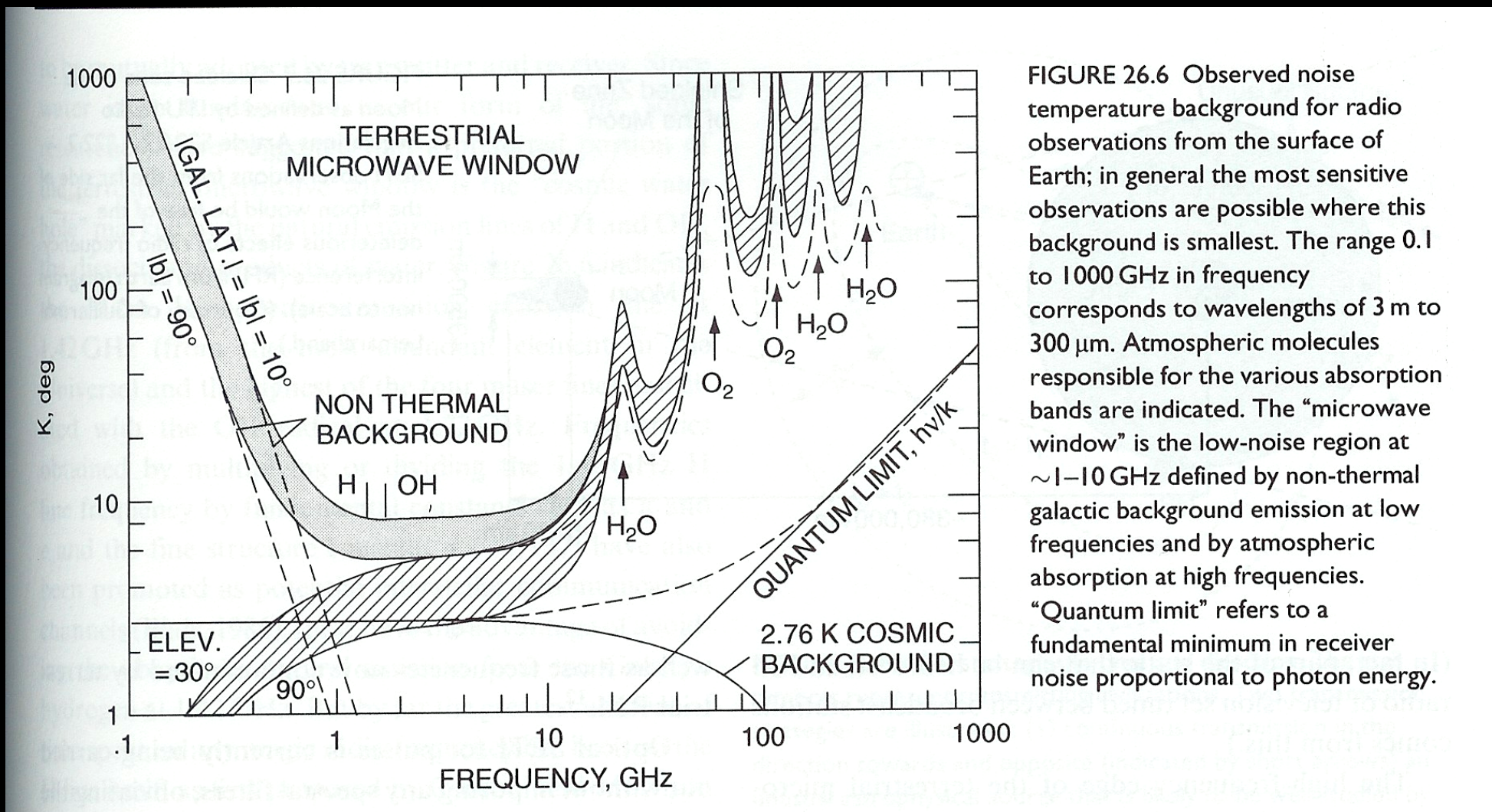


The nearest 10 stars to Earth (excluding our sun, Sol)

1 parsec = 3.2616 light years

| star | absolute visible magnitude | spectral class | luminosity | distance - parsecs |
|--------------------|----------------------------|----------------|------------|--------------------|
| Proxima Centauri C | 15.45 | M5 | | 1.31 |
| Alpha Centauri A | 4.3 | G2 | V | 1.34 |
| Alpha Centauri B | 5.69 | K5 | V | 1.34 |
| Barnard's Star | 13.25 | M5 | V | 1.81 |
| Wolf 359 | 16.68 | M8 | | 2.33 |
| HD 95735 | 10.49 | M2 | V | 2.49 |
| Sirius A | 1.41 | A1 | V | 2.65 |
| Sirius B | 11.56 | WD | VII | 2.65 |
| UV Ceti A | 15.27 | M5 | | 2.72 |
| UV Ceti B | 15.8 | M6 | | 2.72 |

Which frequency range to probe?



From Earth, the best range is 1-10 GHz ($\lambda=3$ -30cm)

If life is water based, communication will happen in the "water hole" (marked by H and OH, i.e. 1.42 GHz to 1.72 GHz)

Current searches

- 1896 Nicola Tesla suggested that radio could be used to contact extraterrestrial life (and thought to have found signals from Mars)
- 1924 “National Radio Silent Day” in the USA (August 21-23)
- 1959 Cocconi and Morrison, Nature - first SETI paper
- 1960 Frank Drake’s Project OZMA
- 1961 First SETI conference at the Green Bank observatory
- 1971 NASA funds Drake’s study for ‘Project Cyclops’ (10^{10} USD!)
- 1979 UC Berkley launches SERENDIP (in 1986 SERENDIP 2)
- 1992 NASA funds the MOP programme, Congress cancels it a year later
- 1995 The SETI institute is founded with private funds



The Allen Telescope Array
Since 2007 (with UC Berkeley)

The Center for SETI Research

Carl Sagan Center for the Study of
Life in the Universe

Center for Education and Public
Outreach



<https://observations.seti.org/> (deactivated?)

1960

Year: 1960
Observers: Drake, F.D.; Project OZMA,
Site(s): NRAO
Frequency: 1420–1420.4 MHz
Resolution: 100 Hz
Objects: 2 stars
[Show](#)

...

~150 campaigns

...

2010

Year: 2010
Observers: Backus, P.; Williams, Peter K. G.; Bower, G. C.; Allen Telescope Array
Team,
Site(s): Allen Telescope Array
Frequency: 1.42 to 1.72 GHz
Resolution: 0.7 Hz
Objects: Galactic Center
[Show](#)

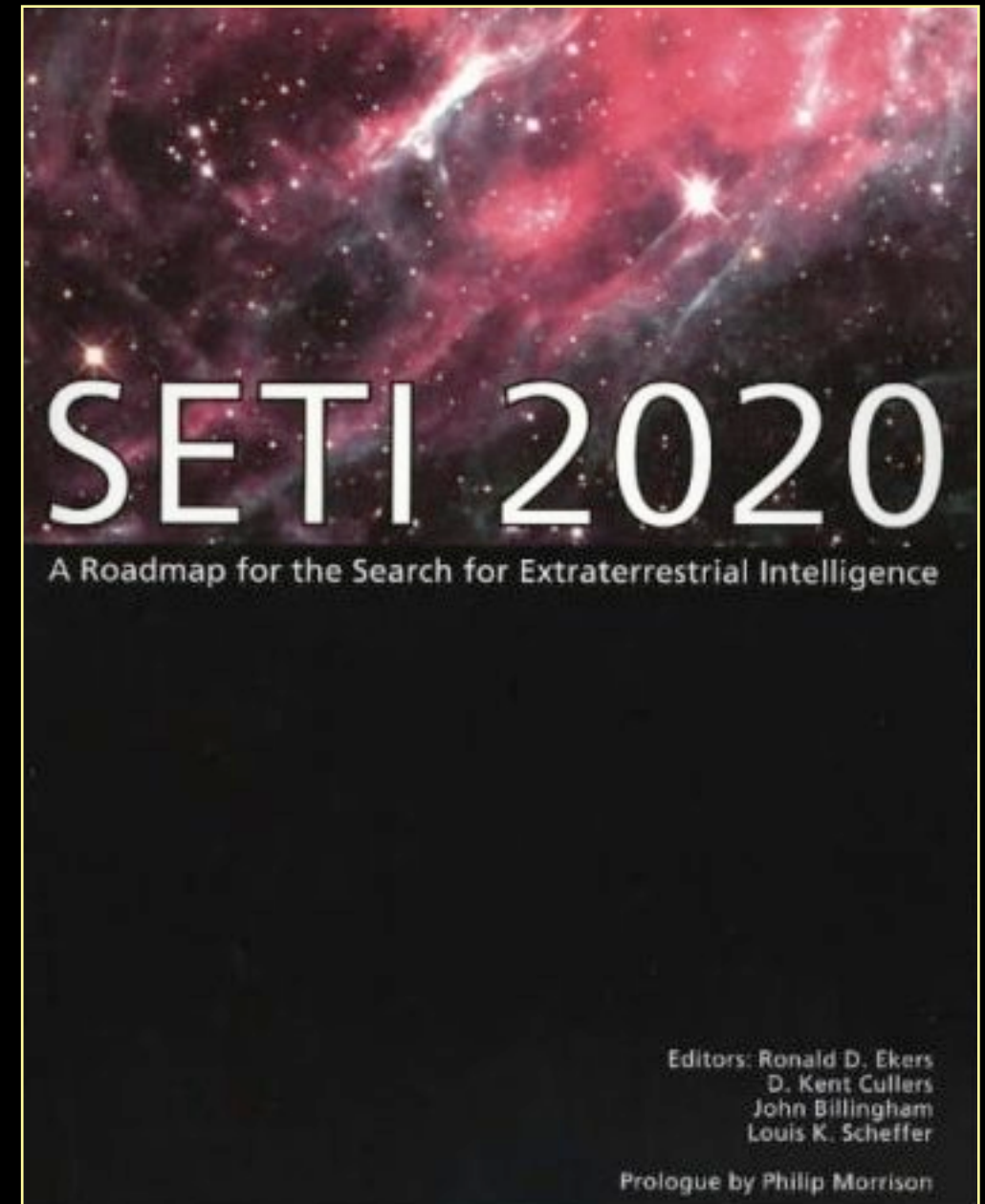
The future of SETI

Harvard: all-sky optical SETI survey
(nanosecond pulses)

Berkeley + SETI: all-sky radio SETI
survey with the Allen Telescope Array
with new funding (100M\$) from the
Breakthrough Listen Initiative

SETI remains marginal science

Search strategies cover only a small
fraction of possible signals...



Lunch break...

Should we communicate?

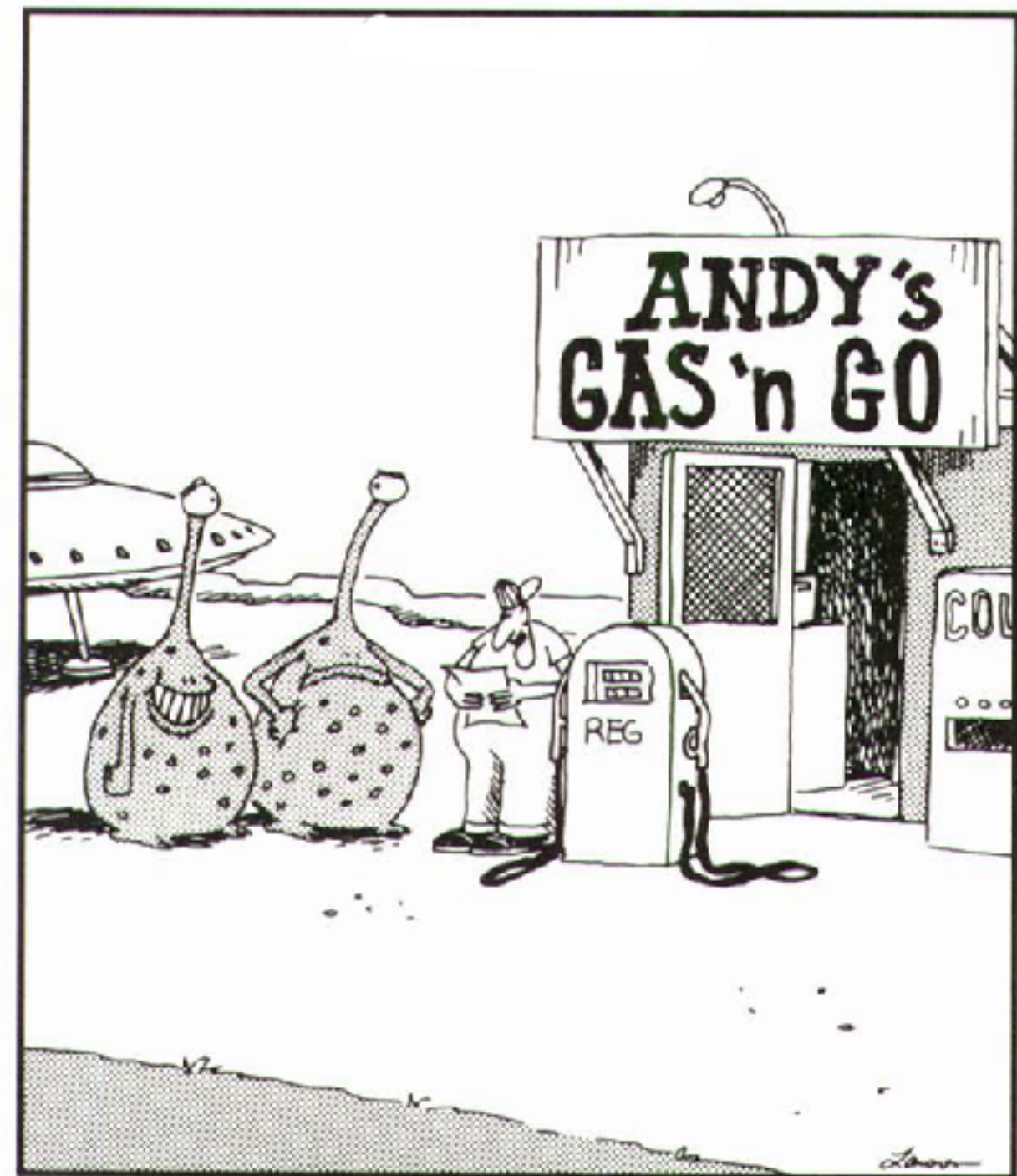
Since the mid 20th century, we are detectable

(Short) broadcast attempts are still taking place

A “Post-detection Protocol” has been adopted by most SETI groups

Current conclusion from a series of workshops (Ekers et al. 2002, ‘SETI 2020’): the strategy should not be to deliberately transmit

Given the Lifespan of a civilization, our transmission would need to last thousands to millions of years - which is beyond the capability of our civilization...



“Shoot! You not only got the wrong planet, you got the wrong solar system. ... I mean, a wrong planet I can understand—but a whole solar system?”

What if we succeed?

Assess the reliability: the Rio Scale (in analogy to the Torino Scale)

| | |
|-------------------------------|--|
| Select Class of Phenomenon: | <div><input type="radio"/> Earth-specific message, or an ET artifact, capable of contact, or a physical encounter</div> <div><input type="radio"/> Omnidirectional message with decipherable information, or a functioning ET artifact or space probe</div> <div><input type="radio"/> Earth-specific beacon to draw our attention, or an ET artifact with a message to mankind</div> <div><input type="radio"/> Omnidirectional beacon designed to draw attention, or an ET artifact with a message of a general character</div> <div><input type="radio"/> Leakage radiation, without possible interpretation, or an ET artifact the purpose of which is understandable</div> <div><input checked="" type="radio"/> Traces of astroengineering, or any indication of technological activity by an extant or extinct civilization at any distance, or an ET artifact, the purpose of which is unknown</div> |
| Select Type of Discovery: | <div><input type="radio"/> SETI/SETA observation; steady phenomenon verifiable by repeated observation or investigation</div> <div><input type="radio"/> Non-SETI/SETA observation; steady phenomenon verifiable by repeated observation or investigation</div> <div><input type="radio"/> SETI/SETA observation; transient phenomenon that has been verified but never repeated</div> <div><input type="radio"/> Non-SETI/SETA observation; transient phenomenon that is reliable but never repeated</div> <div><input checked="" type="radio"/> From archival data; <i>a posteriori</i> discovery without possibility of verification</div> |
| Select Apparent Distance: | <div><input type="radio"/> Within the solar system</div> <div><input type="radio"/> Within a distance which allows communication (at lightspeed) within a human lifetime</div> <div><input type="radio"/> Within the Galaxy</div> <div><input checked="" type="radio"/> Extragalactic</div> |
| Select Credibility of Report: | <div><input type="radio"/> Absolutely reliable, without any doubt</div> <div><input type="radio"/> Very probable, with verification already carried out</div> <div><input type="radio"/> Possible, but should be verified before taken seriously</div> <div><input type="radio"/> Very uncertain, but worthy of verification efforts</div> <div><input checked="" type="radio"/> Obviously fake or fraudulent</div> |

Rio Scale Value

and Importance:

0

None

| Rio | Importance |
|-----|---------------|
| 10 | Extraordinary |
| 9 | Outstanding |
| 8 | Far-reaching |
| 7 | High |
| 6 | Noteworthy |
| 5 | Intermediate |
| 4 | Moderate |
| 3 | Minor |
| 2 | Low |
| 1 | Insignificant |
| 0 | None |

Interpreting Rio Scale Values



Studies show that the impact on world culture and religion will be minor....

Use the Rio Scale calculator to
assess your latest UFO
observation...

[http://avsport.org/IAA/
riocalc.htm](http://avsport.org/IAA/riocalc.htm)

How to go on with
Astrobiology?

Astrobiology

Fundamental questions:

- How does life begin and evolve?
- Does life exist elsewhere in the Universe?
- What is the future of life on Earth?

Each question relies on a complex list of multidisciplinary science questions

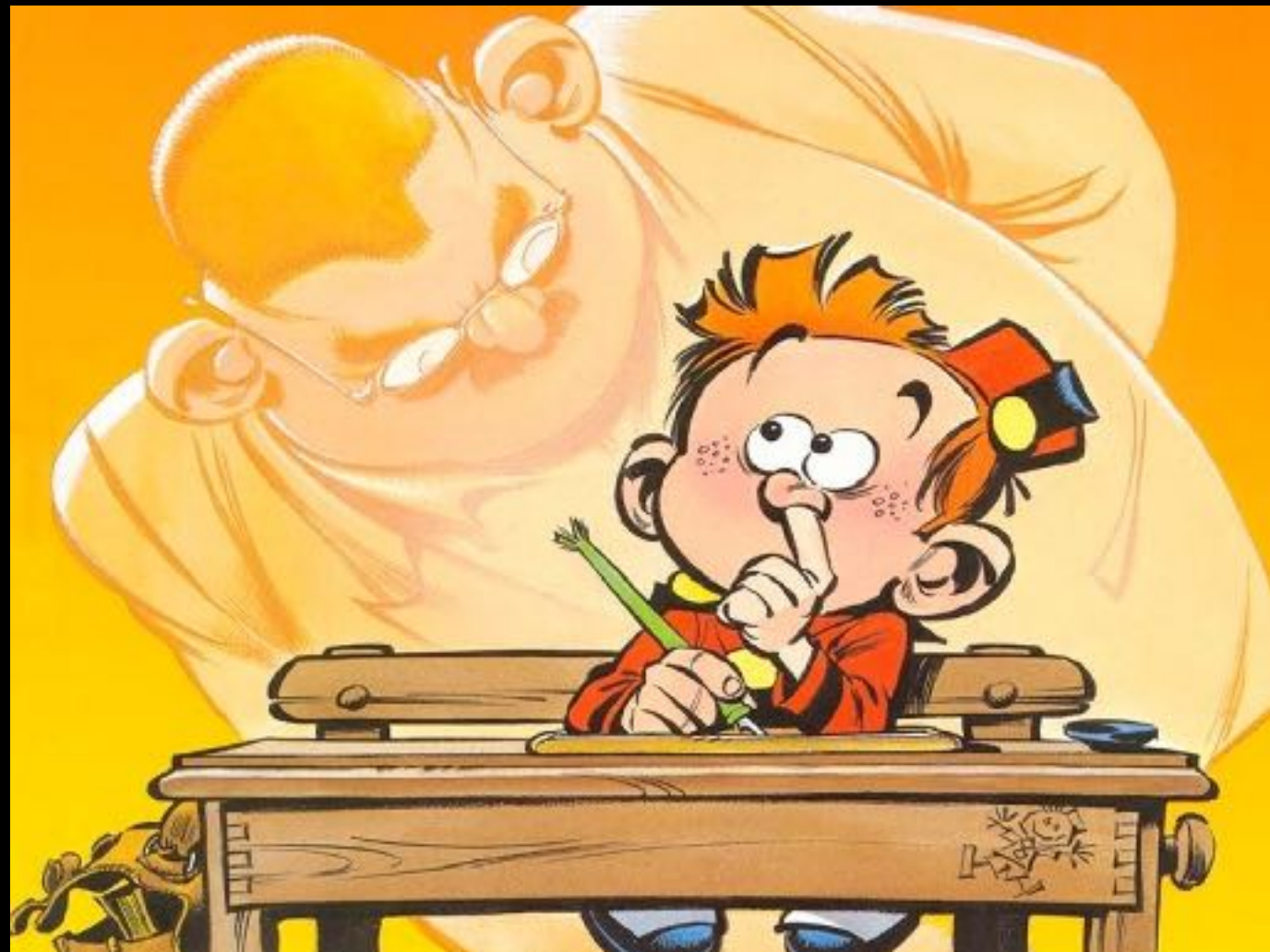


Why create a new *discipline*?

Until the 90s, astrobiological questions were answered in a collaborative framework

Today, the limit is not the scientific knowledge, but the limit of individual disciplines

Astrobiology attempts to break the traditional boundaries of individual disciplines



Societal and Ethical Concerns

- ★ Nuclear reactors and propulsion in space
- ★ Deliberate transportation of organisms between Earth and Mars
- ★ Commercial ventures into space
- ★ Engineering of new lifeforms adapted to live on other worlds
- ★ Terraforming
- ★ Space faring / Colonization

Some of these might not happen in decades or centuries, but it is unforeseeable how fast the field will progress

It is never too early to think about these issues

Scientists will remain the expert decision makers



How can **YOU** contribute to astrobiology?

Do **not** become an *Astrobiologist*

Attend classes in the other disciplines
(Biology, Astronomy, Geology, Chemistry, ...)

Do research on astrobiological questions
relevant to your field
(exoplanets, origin of life, socioeconomics, ...)



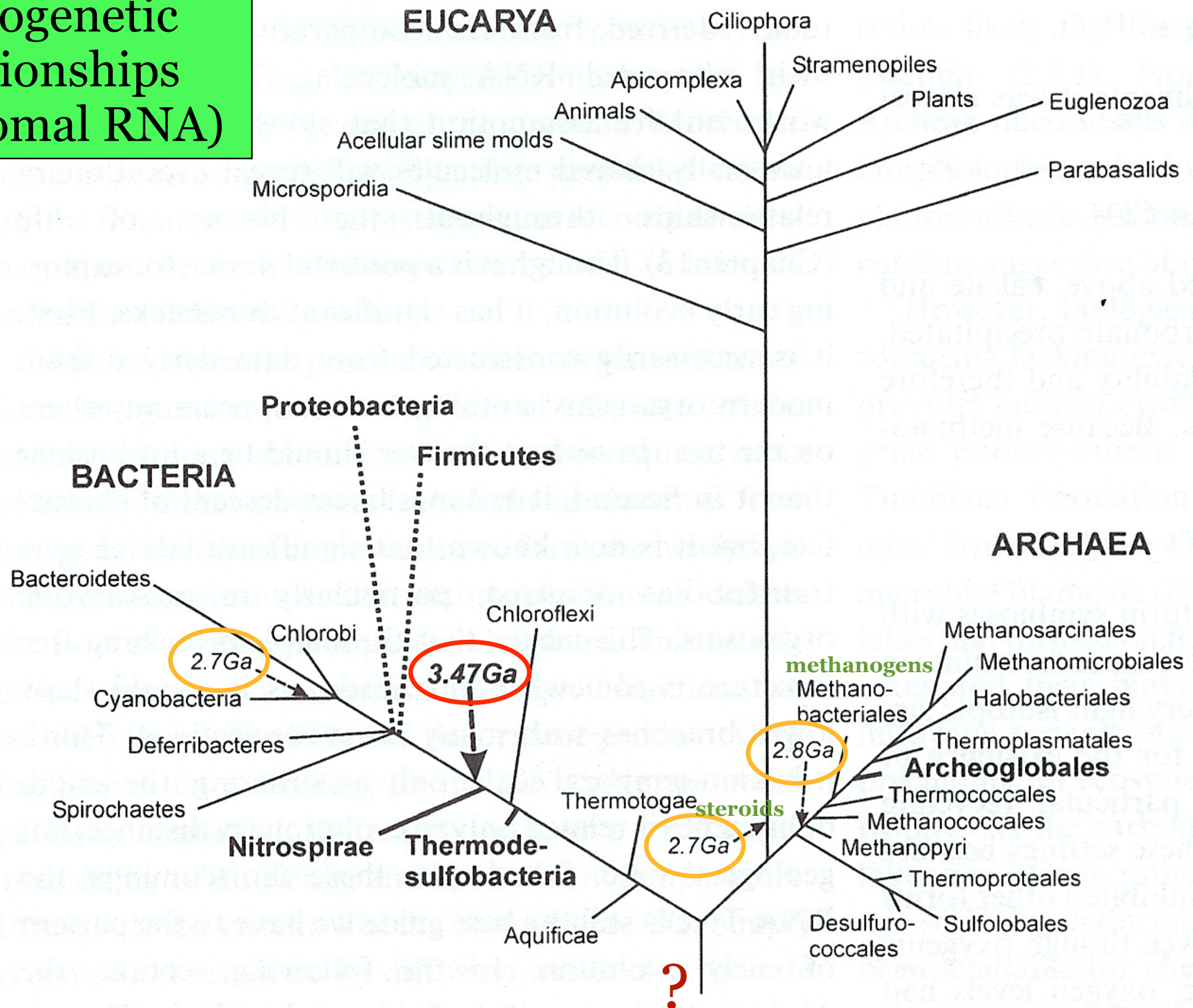
| | | |
|--------------------------|--|---|
| Monday November 20 | Day 1: Definition of Life; Origin of Life; Evolution of Life; Limits of Life 10:00-12:00 & 13:00-14:00 | ✓ |
| Tuesday November 21 | Day 2: Earth Climate History; Mars and Venus Climates 10:00-12:00 & 13:00-14:00 OLD SEMINAR ROOM | ✓ |
| Wednesday November 22 | Day 3: Habitable Places in the Solar System; Mars; Moons of Giant Planets 10:00-12:00 & 13:00-14:00 | ✓ |
| Thursday November 23 | Day 4: Habitable Places beyond the Solar System; Exoplanets properties; Biosignatures 10:00-12:00 & 13:00-14:00 | ✓ |
| Friday November 24 | Day 5: Search for Extraterrestrial Intelligence; Alien Biochemistry 10:00-12:00 & 12:45-13:30 | ✓ |

Overall Summary

Take home ideas from day 1

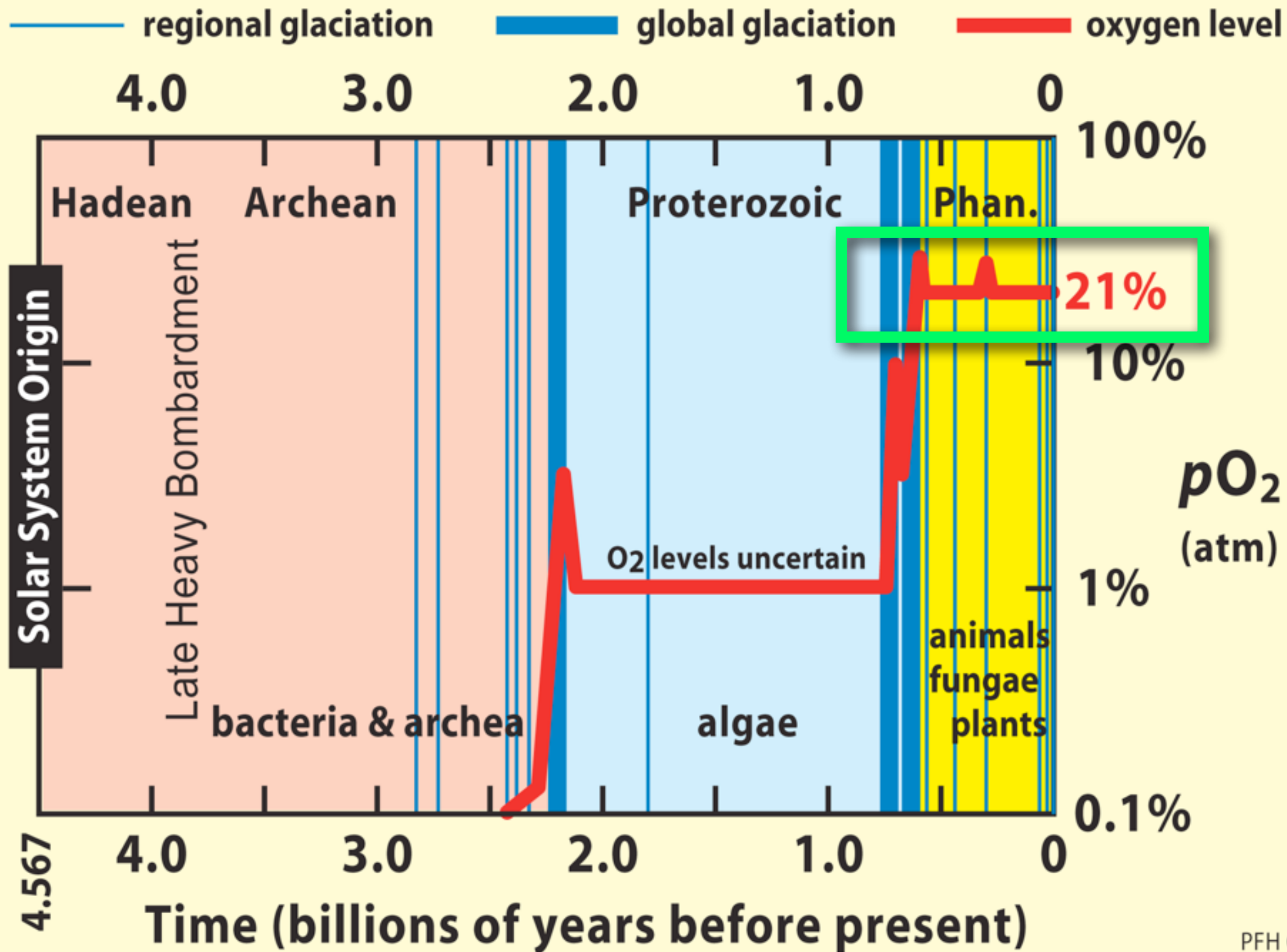
- Defining Life is difficult, as we know only one form of life to date
- Working definition: "A self-sustaining, replicating chemical system, capable of Darwinian evolution."
- Life builds on very common chemistry
- Life appeared on Earth as soon as water was liquid, 4 Gyr ago
- Life evolved and adapted to the most extreme environments

Phylogenetic relationships (ribosomal RNA)



Take home ideas from day 2

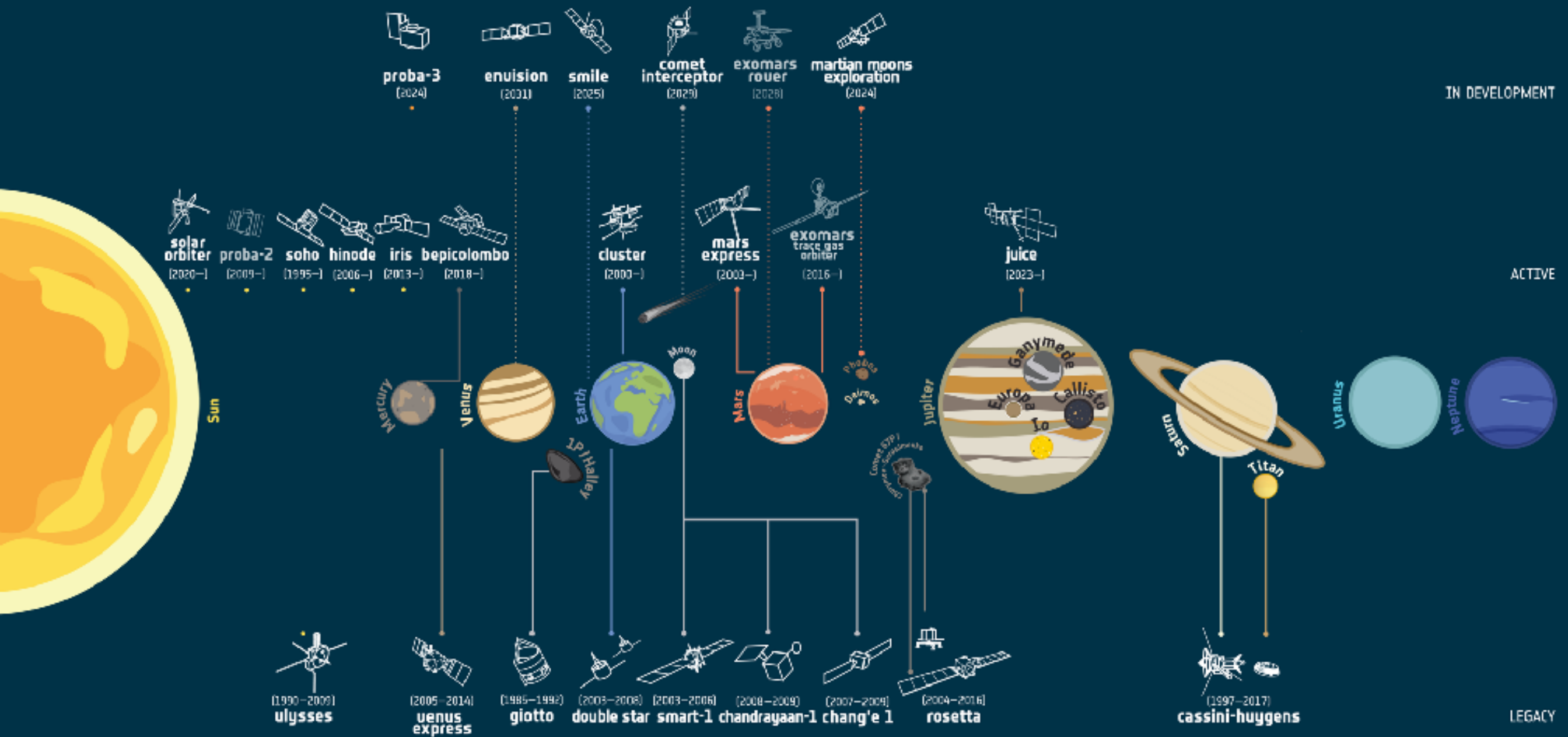
- Solar flux, albedo, greenhouse effect
- Feedback loops drive the climate
- Life modulates the most important volatiles in the atmosphere
- Oxygen appeared with photosynthesis
- Habitability is (currently) linked to liquid water
- Venus experienced a runaway greenhouse effect
- Mars never developed a greenhouse atm

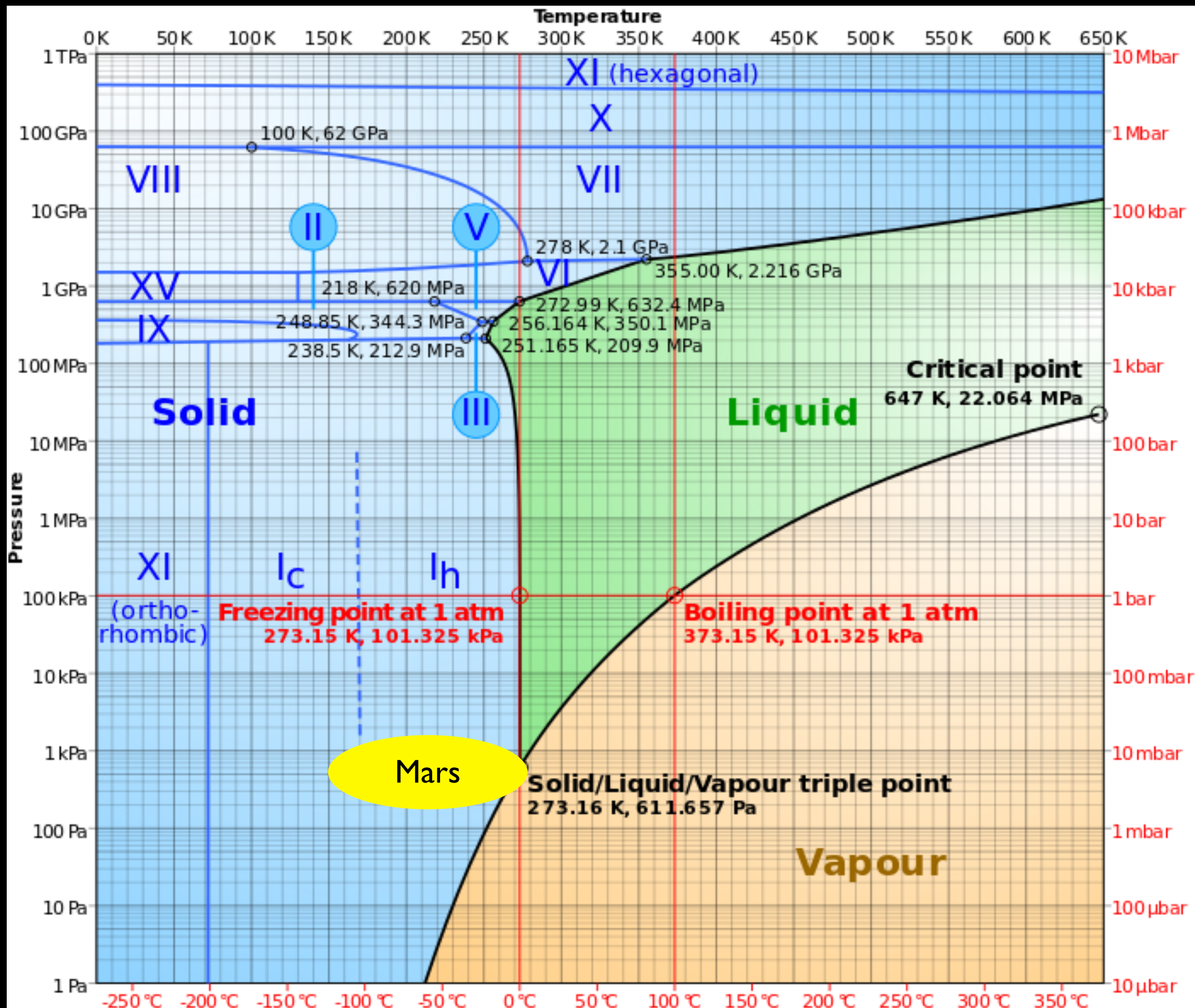


Take home ideas from day 3

- The **temperature gradient** in the initial solar nebulae determines which elements condense where
- The origin of Earth's water is still speculative
- Mars remains the primary target to search for life in the Solar System
- Moons of Jupiter and Saturn host sub-surface liquid water
- Titan would host a weird life form

SOLAR SYSTEM EXPLORERS



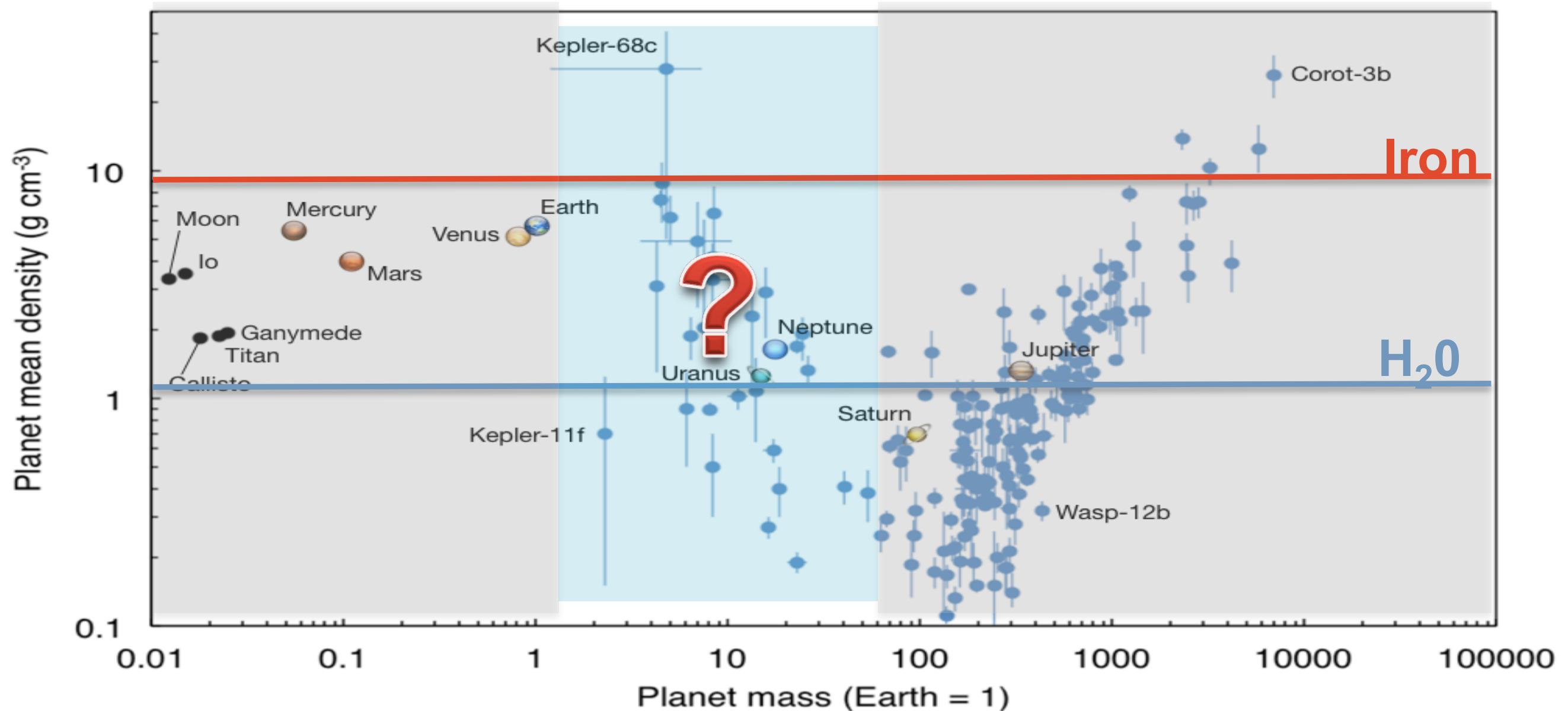


Take home ideas from day 4

- Almost all stars host planets
- Stars $< 1M_{\odot}$ are better suited for life
- Over 5000 exoplanets are known to date, through transiting (radius) and radial velocity methods (mass)
- Hundreds of densities measured
- Spectroscopy of giant planet atmospheres is possible
- Biosignatures in atmospheres are derived from metabolic redox reactions

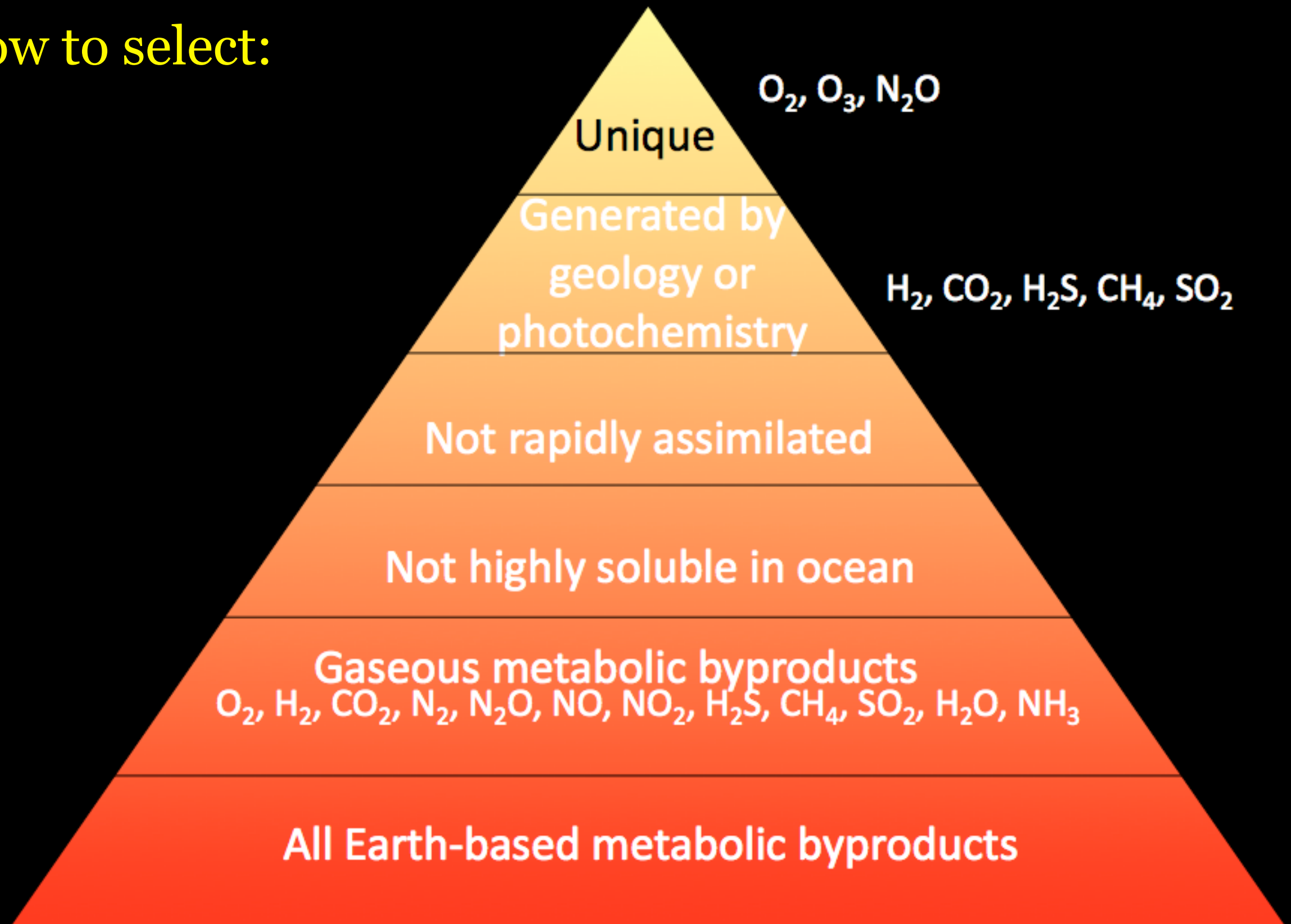
Mass - Radius relation

Mass and Radius allow to derive Density, i.e. Composition



Different atmospheres, different stellar fluxes
could have different biosignatures...

How to select:



Take home ideas from day 5

- Alien biochemistries can expand our idea of life (and habitats)
- Intelligence and consciousness developed with time as part of evolution
- SETI is in fact searching for technology (high risk, high return)
- **Astrobiology:**
 - How does life begin and evolve?
 - Does life exist elsewhere in the Universe?
 - What is the future of life on Earth?

Homework:

- Read some of the many good popular science books on astrobiology and related fields
- Stay curious and open minded

The End for Today

Astrobiology

An Overview

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Thank you!

